

Kinetic Properties of DM-Nitrophen Binding to Calcium and Magnesium

Guido C. Faas,* Kinga Karacs,* Julio L. Vergara,[†] and Istvan Mody*[†]

*Department of Neurology and [†]Department of Physiology, University of California, Los Angeles, California

ABSTRACT Caged-Ca²⁺ compounds such as nitrophenyl-EGTA (NP-EGTA) and DM-nitrophen (DMn) are extremely useful in biological research, but their use in live cells is hampered by cytoplasmic [Mg²⁺]. We determined the properties of Ca²⁺ release from NP-EGTA and DMn by using Oregon green BAPTA-5N to measure changes in [Ca²⁺] after ultraviolet flash photolysis in vitro, with or without Mg²⁺ present. A large fraction (65%) of NP-EGTA, which has a negligible Mg²⁺ affinity, uncages with a time constant of 10.3 ms, resulting in relatively slow increases in [Ca²⁺]. Uncaging of DMn is considerably faster, but DMn has a significant affinity for Mg²⁺ to complicate the uncaging process. With experimentally determined values for the Ca²⁺ and Mg²⁺ binding/unbinding rates of DMn and NP-EGTA, we built a mathematical model to assess the utility of NP-EGTA and DMn in rapid Ca²⁺-uncaging experiments in the presence of Mg²⁺. We discuss the advantages and disadvantages of using each compound under different conditions. To determine the kinetics of Ca²⁺ binding to biologically relevant Ca²⁺ buffers, such as Ca²⁺-binding proteins, the use of DMn is preferable even in the presence of Mg²⁺.

INTRODUCTION

Ca²⁺ is an important second messenger for a variety of physiological and biochemical processes, including muscle contraction, exocytosis and ion-channel gating. Over the last 20 years, a variety of caged-Ca²⁺ compounds have been developed (for review, see Ellis-Davies, 2003) that strongly bind Ca²⁺ and which, upon illumination with a certain wavelength, photolyze into photoproducts (PPs) with much lower affinities for Ca²⁺. Thus controlled by locally applied flashes of light, Ca²⁺ ions become “released” from a “cage”, causing rapid rises in [Ca²⁺]. Of the three most commonly used commercially available caged-Ca²⁺ compounds (DM-nitrophen (DMn), nitrophenyl-EGTA (NP-EGTA), and nitr-5), DMn seems to be the best choice to release caged Ca²⁺ when rapid and large increases in [Ca²⁺] are desired (Nagerl et al., 2000). Compared to nitr-5 and NP-EGTA, DMn has the highest affinity for Ca²⁺ (K_d of 5 nM vs. 145 nM and 80 nM, respectively). In addition, its PPs have the lowest affinity for Ca²⁺ after photolysis (K_d of 3 mM vs. 6.3 μ M and 1 mM for nitr-5 and NP-EGTA, respectively) (Adams et al., 1988; Ellis-Davies, 2003; Ellis-Davies and Kaplan, 1994; Kaplan and Ellis-Davies, 1988). Consequently, ~67–95% of DMn is bound to Ca²⁺ under the normal resting intracellular conditions for most cells (10–100 nM [Ca²⁺]), and most of the bound Ca²⁺ can be released by photolysis. DMn’s relatively high quantum efficiency of 0.18 (Kaplan and Ellis-Davies, 1988) and fast fragmentation into PPs (photolysis of 1.1×10^4 – 8×10^4 s⁻¹ (Ellis-Davies et al., 1996)) make it an almost ideal caged-Ca²⁺ compound. However, its main drawback for physiological experiments involving Ca²⁺ release inside cells is its high affinity for

Mg²⁺ ions normally present in the cytoplasm (Romani and Scarpa, 2000). Large steps in [Ca²⁺] in a Mg²⁺-rich environment can also be accomplished by using NP-EGTA as a Ca²⁺ cage. NP-EGTA has a very low affinity for Mg²⁺ (9 mM) and a relatively large flash-induced change in Ca²⁺ affinity (Ellis-Davies and Kaplan, 1994). When using relatively high resting [Ca²⁺] (>1 μ M), the difference in K_d for Ca²⁺ between DMn and NP-EGTA will make only a small difference in the amount of total uncaged Ca²⁺, since Ca²⁺ is bound to >90% of either compound.

We have previously used a mathematical model to describe Ca²⁺ uncaging from DMn (Escobar et al., 1995, 1997) by a short (20-ns) flash of ultraviolet (UV) light. More recently, a similar model was used to describe the Ca²⁺ transients after fast uncaging of Ca²⁺ and its binding to calbindin D28k, thus allowing the study of the kinetic properties of Ca²⁺ binding to this calcium-binding protein (CBP) (Nagerl et al., 2000). Yet, intracellular CBPs normally operate in the presence of Mg²⁺ and some CBPs may even bind this cation, which could interfere with their Ca²⁺ binding properties (Malmendal et al., 1999; Pechere et al., 1977). To emulate the cytoplasmic conditions, it is imperative to investigate the CBP binding kinetics with both Ca²⁺ and Mg²⁺ present. In such an environment the use of NP-EGTA might be warranted because of its negligible affinity for Mg²⁺ (Ellis-Davies and Kaplan, 1994). However, depending on the desired time course of the Ca²⁺ release, certain experiments might require fast step-like changes in [Ca²⁺]. This might be better accomplished by the use of DMn despite its significant affinity for Mg²⁺ (Kaplan and Ellis-Davies, 1988, Ayer and Zucker, 1999). However, a thorough comparison of the kinetic properties of DMn and NP-EGTA is presently unavailable. Because experiments with DMn become confounding whenever Mg²⁺ is present, most studies have been carried out in the absence of Mg²⁺ (Xu et al., 1997, Neher and Zucker, 1993, Nagerl et al.,

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Address reprint requests to Istvan Mody, Dept. of Neurology, University of California, Los Angeles, Reed Building 710, Los Angeles, CA 90095. Tel.: 310-206-4481; Fax: 310-825-0033; E-mail: mody@ucla.edu.

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2000). Yet, if the properties of Mg^{2+} and Ca^{2+} binding to DMn and its PPs could be well characterized, it should be possible to accurately determine the amount of Ca^{2+} uncaged from DMn in the presence of physiological $[Mg^{2+}]$.

Ayer and Zucker (1999) have sought to determine the Mg^{2+} binding and uncaging properties of DMn by monitoring the $[Ca^{2+}]$ with fura-2 in a solution containing DMn, Ca^{2+} , and Mg^{2+} during and after UV illumination at a time resolution of seconds. In contrast, our goal was to develop a model for fast-timescale “steplike” Ca^{2+} uncaging from DMn and NP-EGTA at resting Ca^{2+} concentrations of $>1 \mu M$. Thus we needed to find a correct mathematical description of the properties of DMn, NP-EGTA, and their PPs and the kinetics of DMn and NP-EGTA uncaging. The model provided as many kinetic constants as possible without overly relying on values determined by earlier studies under experimental conditions vastly different from ours. Our model describes the fast uncaging properties of DMn and NP-EGTA with both Ca^{2+} and Mg^{2+} present, allowing us to make predictions and compare their behavior in a physiological environment. The properties of DMn and NP-EGTA thus determined can be used to study the kinetics of Ca^{2+} binding to CBPs under physiological $[Mg^{2+}]$.

MATERIALS AND METHODS

Solutions

All experiments were done with solutions of 5 mM or 10 mM DMn ((4,5-dimethoxy-2-nitrophenyl)-1,2-diaminoethane-*N,N,N',N'*-tetrasodium salt; Calbiochem, San Diego, CA) or 5 mM NP-EGTA (*o*-nitrophenyl-ethylene glycol-bis(2-aminoethyl ether)-*N,N,N',N'*-tetraacetic acid; Molecular Probes, Eugene, OR) as Ca^{2+} cage with 120 mM KCl and 40 mM HEPES (pH 7.30). Depending on the experiment, a combination of $CaCl_2$ and/or $MgCl_2$ and/or Oregon green BAPTA-5N (OGB-5N, Molecular Probes) was added to these solutions (see below).

Both DMn and NP-EGTA solutions were prepared freshly before every experiment. Fresh stock solutions of DMn and NP-EGTA (15–20 mM) were tested for their accuracy in concentration by measuring their absorbance with a spectrophotometer (Hewlett-Packard 8453, Palo Alto, CA) and using an extinction coefficient of $4.33 \times 10^3 M^{-1} cm^{-1}$ at 350 nm for DMn (Kaplan and Ellis-Davies, 1988) and $5.52 \times 10^3 M^{-1} cm^{-1}$ at 260 nm for NP-EGTA (Ellis-Davies and Kaplan, 1994). For these measurements, small samples were diluted to yield [DMn] or [NP-EGTA] of 50 μM in a solution containing 4 mM EGTA to remove all free Ca^{2+} . We found that most stock solutions contained ~80–90% of the expected DMn or NP-EGTA concentration based on the manufacturer’s specifications (amount of product per vial and molecular weight). To accurately account for the amounts of Ca^{2+} cage, the actual [DMn] and [NP-EGTA] in our experimental solutions were calculated, and these corrected values were used in our analysis. Experiments were carried out at room temperature (~24°C). The experimentation room was outfitted with a yellow filter (500 nm long-pass) for environmental light, to avoid unwanted photolysis. All chemicals were obtained from Sigma-Aldrich (St. Louis, MO) unless otherwise mentioned. Values are expressed as mean \pm SE.

Solution preparation and steady-state $[Ca^{2+}]$ measurements with Ca^{2+} sensitive electrodes

To measure the equilibrium dissociation constants of DMn for Ca^{2+} ($K_{d(CaDMn)}$) and Mg^{2+} ($K_{d(MgDMn)}$), we monitored the free $[Ca^{2+}]$ while

titrating Ca^{2+} -free solutions of DMn (starting with 100 or 200 μl , 5 or 10 mM) with aliquots (1–3 μl) of similar solutions containing 1–100 mM $CaCl_2$. After thoroughly mixing every added aliquot, the $[Ca^{2+}]$ was measured with custom-made Ca^{2+} -sensitive electrodes. This procedure was repeated until a complete binding assay was obtained. The Ca^{2+} -sensitive electrodes were made of a small plastic (PVC) tube (10–20 mm long, diameter 0.7 mm); ~0.5–1 mm on one side of this tube was filled with a 1:1 mixture of 1), 10% (w/w) Ca^{2+} -selective ionophore II (Fluka/Sigma-Aldrich, St. Louis, MO) and 2%) 2% (w/w) Na tetraphenylborate dissolved in 2-nitrophenyloctyl ether, and 2), 8–10% (w/w) polyvinylchloride dissolved in tetrahydrofuran. A thin Ca^{2+} -selective membrane formed after evaporation of the tetrahydrofuran. The electrodes were stored in the dark at room temperature for at least 2 days before use. Calibration solutions consisted of KCl-based pCa standards (Rink et al., 1980) and the pCa 6 standard was used as a reference solution in Ca^{2+} -electrodes. New electrodes were filled with the reference solution at least 12 h before use to allow complete saturation of the membrane. Each electrode was calibrated before every experiment and checked for drift after each experiment. Electrodes were used only if they behaved linearly in the pCa range of 3–7.5 with a slope of at least 27 mV/pCa unit and drifted <1 mV/h. Compensations were made for any drift <1 mV/h. Measurements were accepted if readouts were stable for at least 1 min. The electrodes were insensitive to $[Mg^{2+}]$ up to 10 mM. Since total Mg^{2+} concentrations did not exceed 5 mM in any of our experiments, the Ca^{2+} -sensitive electrodes gave accurate readings of $[Ca^{2+}]$ with Mg^{2+} present. To determine $K_{d(CaDMn)}$ in Mg^{2+} -free solutions we used a simple Scatchard analysis. However, when Mg^{2+} was present, the analysis became more complex. As shown in Appendix A (see Supplementary Material), we were able to express the relative amount of Ca^{2+} -bound DMn ($[CaDMn]/[DMn]_T$) as a function of the measured parameter $[Ca^{2+}]$, of the known parameters $[Ca^{2+}]_T$, $[Mg^{2+}]_T$, and $[DMn]_T$, and of the unknown parameters $K_{d(CaDMn)}$ and $K_{d(MgDMn)}$. The values of $K_{d(CaDMn)}$ and $K_{d(MgDMn)}$ were determined by iterative fitting of the data (Excel, Microsoft, Redmond WA).

Dynamic Ca^{2+} measurements

To measure the dynamics of Ca^{2+} uncaging from DMn or NP-EGTA with high temporal resolution, we used UV-flash photolysis of DMn or NP-EGTA. For these uncaging experiments, all solutions contained 100 μM OGB-5N to monitor rapid changes in Ca^{2+} concentrations. The initial free $[Ca^{2+}]$ was titrated to 1–3 μM so that $>99\%$ of the DMn or $\sim 95\%$ of the NP-EGTA had Ca^{2+} bound.

We used a setup comparable to that described earlier (Escobar et al., 1995, 1997; Nagerl et al., 2000). Briefly, it consisted of a small chamber (20 μl) mounted on an inverted microscope equipped for epifluorescence (IM35; Zeiss, Oberkochen, Germany) with a 505-nm dichroic mirror and 510 LP emission filter (Chroma Technology, Rockingham, VT). In the chamber, the polished end of a silica multimode optical fiber (diameter 800 μm ; Thorlabs, Newton NJ) was mounted to deliver 20 ns flashes of UV light (347 nm) from a frequency-doubled ruby laser (Lumonics, Rugby, UK) to photolyze DMn or NP-EGTA. To excite the OGB-5N molecules, an argon laser (model 95, Lexel, Fremont, CA; 488 nm, 1 W) was focused through the epifluorescence illumination port of the microscope with a 20 \times objective (Fluo20, Nikon, Tokyo, Japan), forming a small illumination spot directly in front of the optical fiber. The relatively small spot size of the excitation light (1–10 μm) compared to the large area of UV illumination (cone with a minimal diameter of 800 μm) ensured minimal diffusion effects during the time span (200 ms) of $[Ca^{2+}]$ changes. The $[Ca^{2+}]$ transients changed substantially only when the illumination spot was moved toward the outer edge of the optical fiber, i.e., to the edge where uncaging took place (not shown). This indicated that diffusion artifacts on the timescale of our measurements occur only very close to the edge of the UV-illuminated area, but not in the area of data collection. The fluorescence of OGB-5N was measured with a photodiode (PIN-HR008, UDT Sensors, Hawthorne, CA) in the focal plane of the microscope. The small diameter of the photodiode (200 μm) minimized errors caused by diffusion in the z axis. Despite the appropriate optical filters and the peak

excitation spectrum of OGB-5N at 494 nm, the high-energy UV flashes still induced brief but large optical transients that saturated the detection system. To avoid these artifacts, a patch-clamp amplifier (Axopatch 200A, Axon Instruments, Foster City, CA) with an integrating headstage was used to measure the currents generated by the photodiode. The feedback capacitor of the headstage was short-circuited (reset) and its readout was blanked exactly at the time of the UV flash so that there was no signal measured at the instant of the flash. The analog signal was low-pass filtered at 10 kHz with the 8-pole Bessel filter of the amplifier, digitized at 50 kHz (PCIO-MIO-16XE-10, National Instruments, Austin, TX), and sampled on a PC with a custom-made program (EVAN) written in LabView (National Instruments). A pulse generator (4000 PG, Neuro-Data Instruments, New York, NY) was used to trigger the sampling, UV laser, headstage reset and blanking, and the shutter for the OGB-5N excitation light.

For a typical experiment, a 10–15 μl droplet of DMn solution was placed in the recording chamber. During each flash, only 0.05–1.5% of the DMn was uncaged in the spot directly in front of the optical fiber (predicted by our model, see Results). However, to avoid significant changes in baseline conditions due to excessive uncaging of DMn or evaporation, no more than three flash-evoked transients were acquired from each droplet. Measurements in the same droplet were performed at least 1 min apart to ensure that all the components in the droplet returned to steady-state baseline conditions. A peak detection circuit located in the laser head measured the energy in each UV flash. The fluorescence transients were stored for offline analysis with a computer model (see below).

Absorbance spectra of the DMn PPs

The absorbance spectra of the PPs formed from DMn, Ca^{2+} -bound DMn and Mg^{2+} -bound DMn were measured with a spectrophotometer (Hewlett-Packard 8453). To produce the various PPs, the following three solutions were exposed for 5 min to the light of a Hg arc-lamp: 50 μM DMn with 10 mM EDTA, 50 μM DMn with 8 mM CaCl_2 , and 50 μM DMn with 8 mM MgCl_2 . To ensure that all PPs were free of Ca^{2+} or Mg^{2+} , the solutions were mixed 1:1 with a 10-mM EDTA solution. The absorption spectra were measured directly after the exposure and mixing with the EDTA solution.

Properties of OGB-5N

We used the low-affinity dye OGB-5N because of its fast kinetics of Ca^{2+} binding and unbinding needed for tracking the expected rapid changes in $[\text{Ca}^{2+}]$. To determine the properties of OGB-5N, we measured its fluorescence in drops containing 5 μM OGB-5N dissolved in the pCa standards. For one batch of the dye (Molecular Probes, lot 34B1-2) we found a K_d of 29.3 μM and an $F_{\text{max}}/F_{\text{min}}$ ratio of 10:8. For another batch (lot 15C1-2) we measured a K_d of 36.1 μM and an $F_{\text{max}}/F_{\text{min}}$ ratio of 40:0. We have no explanation for the variability between the batches other than the fact that specific contaminations might occur in different batches, as stated by the supplier (Molecular Probes). We tested fluorescence changes as a function of Mg^{2+} concentration and found that OGB-5N had no affinity for Mg^{2+} in the concentration range used (<10 mM). To determine the rate constants of Ca^{2+} binding to OGB-5N, an uncaging experiment was done with a DMn solution that had a free $[\text{Ca}^{2+}]$ of only ~ 2 nM. In this case, no more than 30% of the DMn will bind Ca^{2+} , and consequently any uncaged Ca^{2+} will rapidly rebind to free DMn. Compared to Ca^{2+} dye kinetics, the Ca^{2+} signal generated by uncaging from and rebinding to DMn at low free $[\text{Ca}^{2+}]$, will approach a delta function (Escobar et al., 1997). Therefore, with low $[\text{Ca}^{2+}]$, the decay of $\Delta F/F$ will resemble the unbinding of Ca^{2+} from OGB-5N. Indeed the signal could be accurately fitted with a single decay time constant of 133 μs for lot 34B1-2 and 115 μs for lot 15C1-2. These numbers indicate a $k_{\text{off(D)}}$ of $7.52 \times 10^3 \text{ s}^{-1}$ and a $k_{\text{on(D)}}$ of $2.6 \times 10^8 \text{ M}^{-1} \text{ s}^{-1}$ for lot 34B1-2 and a $k_{\text{off(D)}}$ of $8.67 \times 10^3 \text{ s}^{-1}$ and a $k_{\text{on(D)}}$ of $2.4 \times 10^8 \text{ M}^{-1} \text{ s}^{-1}$ for lot 15C1-2. These values were used in the mathematical model (see below) to describe the properties of the two batches of OGB-5N used in the various experiments.

Data analysis

All data were analyzed using MS Excel (Microsoft) and Berkeley Madonna 8.0 (University of California, Berkeley, CA). To determine the kinetic parameters (association and dissociation rates) from the fluorescence recordings we used a mathematical model incorporating all of the reactions in the uncaging solution (see Supplementary Material, Appendix B).

The general-purpose differential equation solver Berkeley Madonna was used to fit the simulations to the fluorescence recordings by iterating the parameters with the fourth-order Runge-Kutta method. First, all data were expressed as

$$\frac{\Delta F(t)}{F} = \frac{(F(t) - F_{\text{rest}})}{F_{\text{rest}}} \quad (1)$$

where $F(t)$ is the fluorescence acquired over time t , with $t = 0$ at the time of the flash, and F_{rest} is the resting fluorescence averaged over 50 ms before flash delivery. To more accurately fit the fast rising phase while avoiding bias from late slow-decaying phase of the fluorescence transients, data points were omitted exponentially in time from the fitting routine.

To increase the accuracy of the fits to the model, sets of 10 traces obtained at different flash energies were fitted simultaneously. For each set of traces, all fitted parameters except for the amount of uncaging, which is dependent on the flash energy, were constrained to be equal between the individual traces.

RESULTS

Ca^{2+} release from NP-EGTA

Due to its very low Mg^{2+} affinity ($K_{\text{d}(\text{MgNP-EGTA})} = 9 \text{ mM}$) and relatively large quantum yield (0.23 vs. 0.18 for DMn) (Ellis-Davies et al., 1996; Ellis-Davies and Kaplan, 1994), NP-EGTA has been considered the ‘‘caged- Ca^{2+} ’’ molecule of choice whenever the experimental conditions demand the presence of physiological $[\text{Mg}^{2+}]$. However, careful evaluation of the kinetic properties of Ca^{2+} binding to NP-EGTA and its photoproducts after photolysis is necessary to fully assess the potential of this probe as a tool for the rapid control of the $[\text{Ca}^{2+}]$ inside a cell. Fig. 1 shows that immediately after uncaging NP-EGTA, a short spike in $[\text{Ca}^{2+}]$ was observed, followed by a much slower rise in the $[\text{Ca}^{2+}]$, ultimately leading to a lasting steady-state level. To analyze quantitatively the kinetics of the fluorescence traces and to determine the properties of NP-EGTA, we used a mathematical model of the uncaging and the various equilibrium reactions (see Appendix B, Supplementary Material). This model is extensively described below for DMn. Sets of NP-EGTA uncaging traces could only be accurately fit when using two uncaging time constants (Fig. 1) and a dissociation constant ($K_{\text{d}(\text{CaNP-EGTA})}$) of $50 \pm 4 \text{ nM}$ that changes into a K_d of $11 \pm 1 \text{ mM}$ for the photoproducts of NP-EGTA ($n =$ four sets of 10 traces; for detailed rate constants, see 1). We found that $35 \pm 4\%$ of NP-EGTA uncaged with a time constant, τ_{fast} , of $2.8 \pm 0.8 \mu\text{s}$, whereas $65 \pm 4\%$ uncaged with a τ_{slow} of $10.3 \pm 0.7 \text{ ms}$. The Ca^{2+} released from the rapidly uncaging fraction of NP-EGTA is quickly rebound to Ca^{2+} -free NP-EGTA, whereas 2/3 of NP-EGTA uncages slowly, causing a gradual increase in $[\text{Ca}^{2+}]$. The slow Ca^{2+} release from NP-EGTA may limit its use in physiological experiments

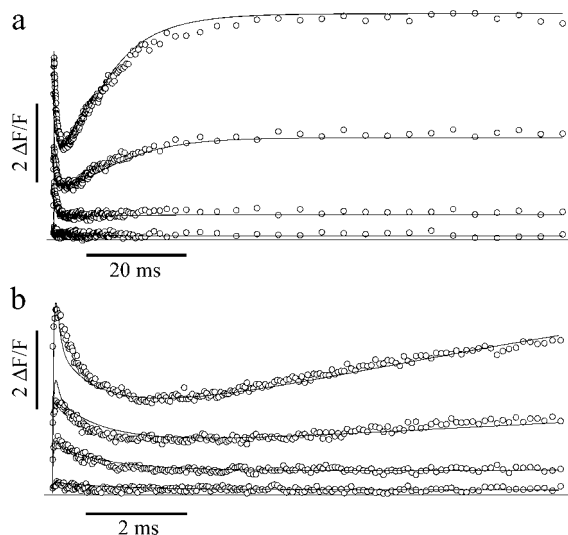
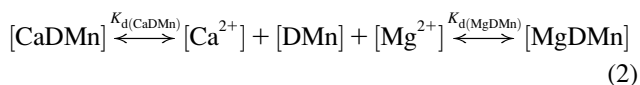


FIGURE 1 Fluorescence data of Ca^{2+} uncaging from NP-EGTA (open circles) using different flash energies. At a resting $[\text{Ca}^{2+}]$ of $2.1 \mu\text{M}$, NP-EGTA causes a large biphasic change in $\Delta F/F$ and thus in $[\text{Ca}^{2+}]$ (a and b). The initial rise in $[\text{Ca}^{2+}]$ is more clearly depicted on a faster timescale in b. Ca^{2+} uncaging from NP-EGTA was fitted simultaneously to sets of 10 traces evoked by different flash intensities (see text for details), with the models using two uncaging time constants (solid lines in a; only four traces are shown for clarity).

despite its low Mg^{2+} affinity. Conversely, DMn might have beneficial kinetic properties over those of NP-EGTA even with Mg^{2+} present. To compare the two caged compounds, we measured the binding constants and uncaging kinetics of Ca^{2+} from DMn under physiological $[\text{Mg}^{2+}]$.

Steady-state analysis of Ca^{2+} and Mg^{2+} binding to DMn

First we determined the steady-state $K_{d(\text{CaDMn})}$ by titrating a DMn (5 mM) solution with Ca^{2+} while measuring the $[\text{Ca}^{2+}]$. The $K_{d(\text{CaDMn})}$ was determined by fitting the data in Scatchard plots as shown in Fig. 2 a (solid circles). The $K_{d(\text{CaDMn})}$ determined in this manner was $8.1 \pm 1.4 \text{ nM}$ ($n = 13$), comparable to values found previously using a variety of techniques (Escobar et al., 1997; Kaplan and Ellis-Davies, 1988). The same titrations with Ca^{2+} were made to determine the $K_{d(\text{MgDMn})}$, but with a constant amount of Mg^{2+} (1–5 mM) present during the titration. The Mg^{2+} ions interfered with Ca^{2+} binding according to the following scheme:



In the presence of Mg^{2+} , it took relatively higher concentrations of free Ca^{2+} to occupy the same number of binding sites on DMn. This is illustrated in Fig. 2 a, where the data are plotted as Scatchard plots, and in Fig. 2 b, where some of the data of Fig. 2 a, are plotted as a steady-state

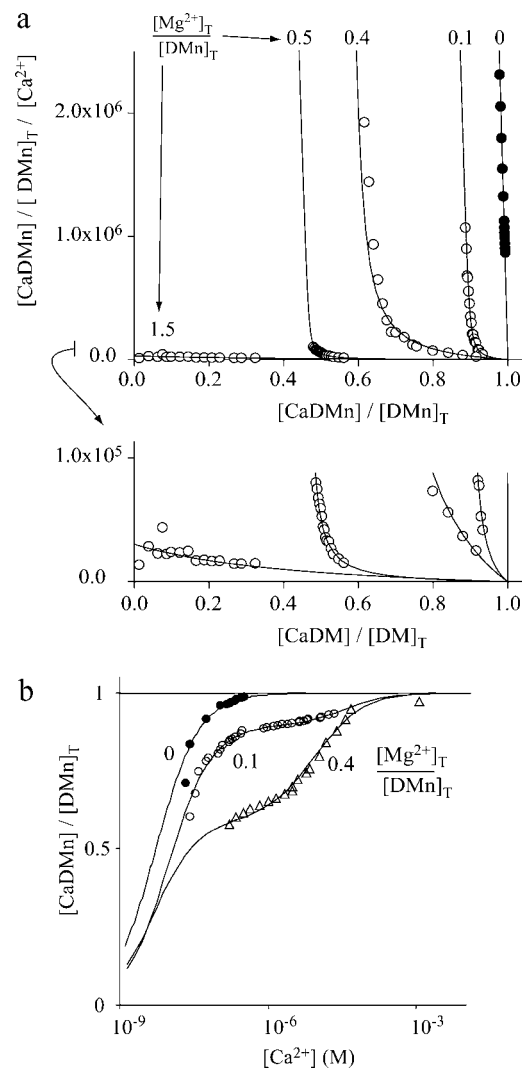


FIGURE 2 Steady-state Ca^{2+} - and Mg^{2+} -binding properties of DMn. (a) Scatchard plots of the titration with Ca^{2+} of typical solutions containing DMn without Mg^{2+} present ($[\text{Mg}^{2+}]_{\text{T}}/[\text{DMn}]_{\text{T}} = 0$; ●) and with Mg^{2+} present ($[\text{Mg}^{2+}]_{\text{T}}/[\text{DMn}]_{\text{T}} = 0.1, 0.4, 0.5, \text{ or } 1.5$; ○). A Ca^{2+} -sensitive electrode was used to measure the free $[\text{Ca}^{2+}]$ and the $[\text{CaDMn}]$ was calculated as the difference between added and free $[\text{Ca}^{2+}]$. Linear fit through the solid circles was made to determine the $K_{d(\text{CaDMn})}$ ($1/\text{slope}$). The open circles represent data from titrations with the indicated ratios of Mg^{2+} present relative to the amount of DMn. The open circles were fitted with a function that described the relationship between $[\text{CaDMn}]$ and $[\text{Ca}^{2+}]$ as a function of $[\text{Ca}^{2+}]_{\text{T}}$, $[\text{Mg}^{2+}]_{\text{T}}$, $[\text{DMn}]_{\text{T}}$, $K_{d(\text{CaDMn})}$, and $K_{d(\text{MgDMn})}$ to determine $K_{d(\text{CaDMn})}$ and $K_{d(\text{MgDMn})}$ (see Supplementary Material, Appendix A). The inset expands the y axis 10-fold. (b) Some of the data in a are shown as binding curves to clarify that with Mg^{2+} present more Ca^{2+} is needed to occupy all the binding sites of DMn. The binding curves for Ca^{2+} to DMn are shown at three different ratios of $[\text{Mg}^{2+}]_{\text{T}}/[\text{DMn}]_{\text{T}}$: 0, i.e., no Mg^{2+} present (●); 0.1 (○); and 0.4 (△).

binding curve of Ca^{2+} to DMn. With no Mg^{2+} present, the Ca^{2+} binding to DMn can be described by a linear function in the Scatchard plot or a sigmoidal function in the steady-state binding plots (fits through solid symbols in Fig. 2, a and b). When Mg^{2+} is present, the Scatchard plots shift to the

left. This shift becomes more pronounced with increasing $[Mg^{2+}]$ (open circles). The left shift of the Scatchard plots is equivalent to a shift to the right in the steady-state binding curves (Fig. 2 *b*). Thus, as expected, increasing $[Mg^{2+}]$ increased the amount of Ca^{2+} needed to fully occupy all DMn with Ca^{2+} . Furthermore, when Mg^{2+} is present, the Ca^{2+} binding curves cannot be described by a simple sigmoidal function (Fig. 2 *b*) and, similarly, the Scatchard plots can no longer be described by a simple line (Fig. 2 *a*). From the general steady-state equations for the equilibrium condition, as shown in Eq. 2, we derived a function describing the relative amount of CaDMn ($[CaDMn]/[DMn]_T$) as a function of the measured $[Ca^{2+}]$, known parameters such as $[Ca^{2+}]_T$, $[Mg^{2+}]_T$, and $[DMn]_T$, and of unknown parameters $K_{d(CaDMn)}$ and $K_{d(MgDMn)}$ (see Supplementary Material, Appendix A). To determine $K_{d(CaDMn)}$ and $K_{d(MgDMn)}$, this function was fitted to the Scatchard plots (shown in Fig. 2,

a and *b*, as lines through the open symbols). The fitted $K_{d(CaDMn)}$ was 7.0 ± 1.4 nM ($n = 10$), comparable to the $K_{d(CaDMn)}$ found when no Mg^{2+} was present. This indicates that Mg^{2+} ions do not change the affinity of DMn for Ca^{2+} , but merely compete with Ca^{2+} for binding to DMn. For Mg^{2+} binding to DMn we found $K_{d(CaDMn)}$ of 1.5 ± 0.7 μ M.

Kinetic properties for the interaction between Ca^{2+} and DMn

To establish the kinetic properties of Ca^{2+} and Mg^{2+} binding to DMn, we measured the $[Ca^{2+}]$ directly after uncaging DMn in different solutions. The open circles in Fig. 3, *b* and *e*, show changes in OGB-5N fluorescence immediately after photolysis of DMn in the presence of Ca^{2+} ($[Mg^{2+}] = 0$; resting $[Ca^{2+}] = 1\text{--}3$ μ M) at different flash energies of the UV laser. In response to the UV flash,

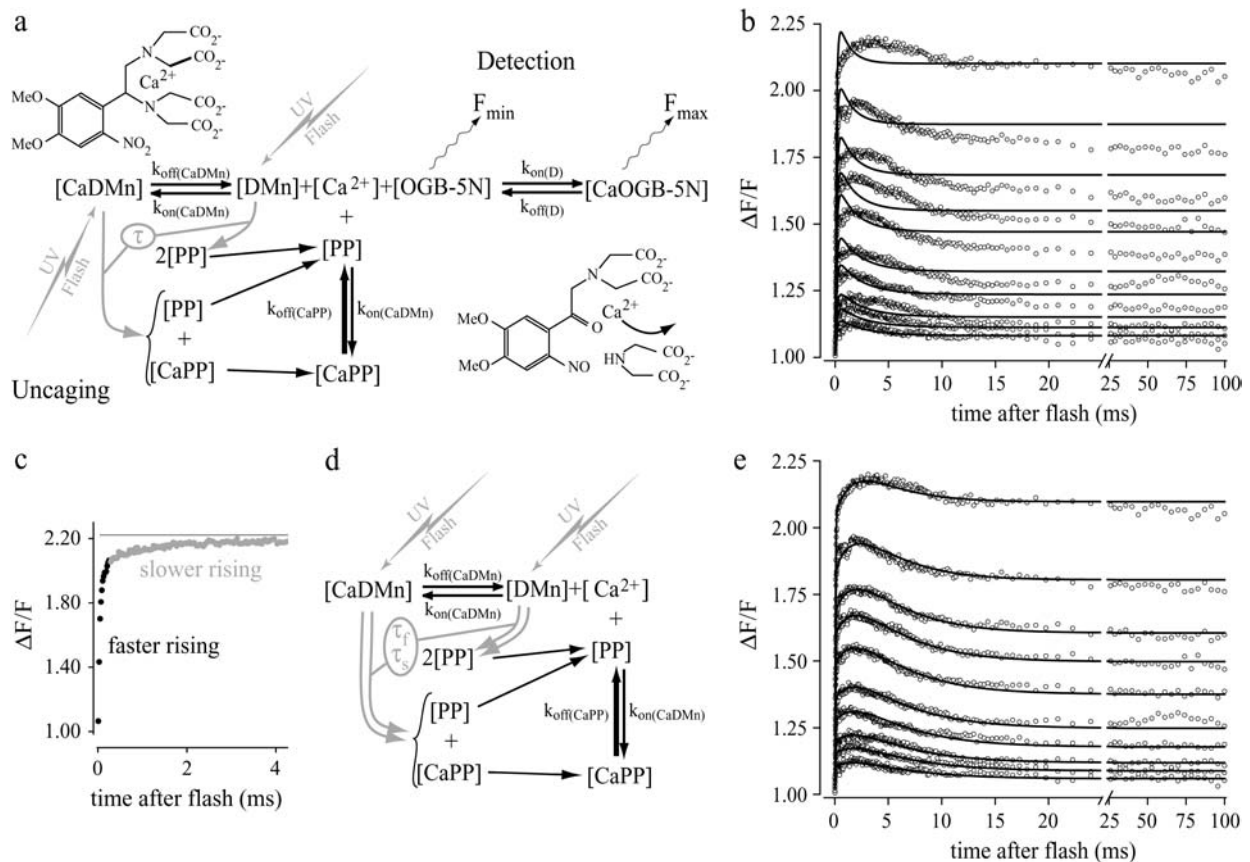


FIGURE 3 Modeling and fitting of Ca^{2+} uncaging from DMn. (a) Scheme of the model to simulate the dynamic redistribution of Ca^{2+} generated by flash photolysis of DMn. Upon flash delivery, a fraction of the DMn uncages (shaded arrows) with a single time constant (τ) into PPs. Free Ca^{2+} levels are detected by the change of OGB-5N fluorescence (F_{min} versus F_{max}) as it binds Ca^{2+} . Insets show structure of DMn and suggested structure of PPs. (b) Typical set of fluorescence data of Ca^{2+} uncaging from DMn (open circles) with different flash energies. Attempts to fit these data with the model shown in *a* failed (solid lines). (c) The initial part of the largest fluorescence transient in *b* is shown on an expanded timescale to demonstrate the two apparent uncaging time constants: a fast one (black) and a slow one (shaded). (d) A part of the scheme in *a* is shown incorporating into the model the two uncaging time constants (τ_f and τ_s , shown in gray). (e) The model with the two uncaging time constants accurately fits the same set of data as in *b*. Note that in the fit to the transients in both *b* and *e* a more accurate fit over the fast rising and falling phases at the start of the traces was ensured by exponentially weighting the fits by omitting data points from fitting toward the end of the sweeps.

a fraction of the total DMn breaks up, releasing Ca^{2+} . Before the flash, CaDMn and DMn, and CaOGB-5N and OGB-5N, are at equilibrium. Upon flash delivery, a fraction of both the Ca^{2+} -free and Ca^{2+} -bound DMn break up to form a PP and Ca^{2+} -bound PP (CaPP), as indicated in Fig. 3 *a*. Since the affinity of PP for Ca^{2+} is much lower than that of DMn, a certain amount of the formed CaPP complexes will dissociate into PP and Ca^{2+} , resulting in a sharp increase in the $[\text{Ca}^{2+}]$. It has been shown (Escobar et al., 1997) that at a high preflash $[\text{Ca}^{2+}]$ photolysis of DMn will lead to a steplike increase in $[\text{Ca}^{2+}]$ since most of the DMn is Ca^{2+} -bound and rebinding of uncaged Ca^{2+} to free DMn molecules becomes negligible. In our experiments, the preflash $[\text{Ca}^{2+}]$ was between 1 and 3 μM , yielding 99.2–99.7% of DMn Ca^{2+} -bound, considering a $K_{\text{d}(\text{CaDMn})}$ of 7–8 nM. Although our experiments are consistent with roughly a square step-shaped rise in $[\text{Ca}^{2+}]$, under these conditions some Ca^{2+} still rebinds to free DMn molecules, leading to a small drop in $[\text{Ca}^{2+}]$ after the initial rise (Fig. 3, *b* and *e*).

To analyze the kinetics of the fluorescence traces and to obtain the properties of DMn, we used a mathematical model describing three equilibrium reactions (Fig. 3 *a*, and Supplementary Material, Appendix B). We first assumed that the uncaging of affected DMn, i.e., the transition of DMn to PP, is governed by a single time constant (τ , Fig. 3 *a*, and Supplementary Material, Appendix B). Furthermore, we assumed that the association rates of Ca^{2+} ($k_{\text{on}(\text{CaDMn})}$) with both DMn and PP were equal. Sets of 10 sweeps obtained at different flash energies were fitted simultaneously to determine $K_{\text{d}(\text{CaDMn})}$, $k_{\text{on}(\text{CaDMn})}$, and the Ca^{2+} dissociation rate for PP ($k_{\text{off}(\text{CaPP})}$). From these values, the dissociation rate of Ca^{2+} from DMn ($k_{\text{off}(\text{CaDMn})}$) and the K_{d} of PP for Ca^{2+} ($K_{\text{d}(\text{CaPP})}$) were calculated. For a given fitted dataset, all corresponding values were constrained to be equal across the traces obtained with different flash energies. Only the fraction of affected DMn, i.e., the amount of uncaging, was allowed to vary between traces. The model was unable to make accurate fits in every one of the 11 experiments (example in Fig. 3 *b*) even when all kinetic parameters for DMn and PP were fitted (thus could be of any value). A high-resolution plot of the rising phase of the fluorescence

transient shows a steep elevation followed by a slower rise of $[\text{Ca}^{2+}]$ (Fig. 3 *c*). This suggests that the transition from DMn to PP might be governed by two time constants instead of one. This is plausible, as DMn breaks up via two different pathways (Il'ichev et al., 2004; Ellis-Davies et al., 1996).

Accordingly, we adapted the model to accommodate the transition breakup of DMn into PP via two pathways, one with a fast time constant (τ_{fast}) and the other with a slow time constant (τ_{slow}) (Fig. 3 *d*). With the bi-exponential uncaging implemented, the model more accurately described the experimental data fitting $K_{\text{d}(\text{CaDMn})}$, $k_{\text{on}(\text{CaDMn})}$, $k_{\text{off}(\text{CaPP})}$, τ_{fast} , τ_{slow} , and the relative contribution of these time constants (Fig. 3 *e* and 1). We found that $67 \pm 2\%$ of DMn uncaged with a τ_{fast} of $15 \pm 2 \mu\text{s}$, whereas $33 \pm 2\%$ uncaged with a τ_{slow} of $2.0 \pm 0.4 \text{ ms}$. The relative distribution of τ_{fast} and τ_{slow} is comparable to the previously reported distribution of the two decay pathways of DMn (Ellis-Davies et al., 1996; McCray et al., 1992). The $K_{\text{d}(\text{CaDMn})}$ for these fits was $7.0 \pm 1.4 \text{ nM}$, which is similar to the values for $K_{\text{d}(\text{CaDMn})}$ found in the titration experiments. The determined rate constants of $2.9 \pm 0.3 \times 10^7 \text{ M}^{-1} \text{ s}^{-1}$ for $k_{\text{on}(\text{CaDMn})}$ and $0.19 \pm 0.03 \text{ s}^{-1}$ for $k_{\text{off}(\text{CaDMn})}$ are comparable to previously published values, as are the values for the PP, with a $K_{\text{d}(\text{CaPP})}$ of $3.3 \pm 0.6 \text{ mM}$ and a $k_{\text{off}(\text{CaPP})}$ of $77 \pm 8 \times 10^3 \text{ s}^{-1}$ (Ayer and Zucker, 1999; Ellis-Davies et al., 1996; Escobar et al., 1995, 1997; Kaplan and Ellis-Davies, 1988). The values of the K_{d} s and rate constants determined by us (see summary in Table 1) agree well with those obtained by others using different techniques. This, together with the highly accurate fits to the $\Delta F/F$ curves, substantiate that our mathematical model provides a realistic representation of the DMn uncaging process. Therefore, in our subsequent experiments the bi-exponential model was used to describe the uncaging of DMn.

Kinetic properties of the interactions between Mg^{2+} and DMn

To study the kinetics of interactions between DMn and Mg^{2+} we carried out uncaging experiments as described above, but in the presence of 1–5 mM Mg^{2+} . Fig. 4 *a* shows uncaging curves generated with similar uncaging energies

TABLE 1 Properties of DMn and NP-EGTA and their photoproducts determined by titration or with flash photolysis

Interaction	Experiment	<i>n</i>	K_{d} (M)	k_{on} ($\text{M}^{-1} \text{ s}^{-1}$)	k_{off} (s^{-1})	τ_{fast} (s)	Fraction τ_{fast} (%)	τ_{slow} (s)
Ca^{2+} -DMn	Ca^{2+} titration	13	$8.1 \pm 1.4 \times 10^{-9}$					
Ca^{2+} -DMn	Ca^{2+} titration + Mg^{2+}	10	$7.0 \pm 1.4 \times 10^{-9}$					
Ca^{2+} -DMn	Photolysis Ca^{2+}	11	$7.6 \pm 1.2 \times 10^{-9}$	$2.9 \pm 0.3 \times 10^7$	$1.9 \pm 0.2 \times 10^{-1}$	$15 \pm 2 \times 10^{-6}$	67 ± 2	$2.0 \pm 0.4 \times 10^{-3}$
Ca^{2+} -PP1	Photolysis Ca^{2+}	11	$3.1 \pm 0.5 \times 10^{-3}$	2.9×10^7	$6.9 \pm 0.8 \times 10^4$		← Determined with model (a) or (b)	
Ca^{2+} -PP2	Photolysis Ca^{2+} + Mg^{2+}	12	$89 \pm 41 \times 10^{-6}$	2.9×10^7	$3.6 \pm 1.7 \times 10^5$		← Determined with model (b)	
Mg^{2+} -DMn	Ca^{2+} titration + Mg^{2+}	10	$1.5 \pm 0.7 \times 10^{-6}$					
Mg^{2+} -DMn	Photolysis Ca^{2+} + Mg^{2+}	12	$5.0 \pm 0.9 \times 10^{-4}$	$3.3 \pm 1.5 \times 10^4$	$2.4 \pm 1.2 \times 10^1$		Determined with model (a)	
Mg^{2+} -PP	Photolysis Ca^{2+} + Mg^{2+}	12	5.4 ± 3.5	3.3×10^4	$92 \pm 53 \times 10^3$			
Mg^{2+} -DMn	Photolysis Ca^{2+} + Mg^{2+}	12	$1.9 \pm 0.2 \times 10^{-6}$	$1.3 \pm 0.3 \times 10^5$	$2.0 \pm 0.3 \times 10^{-1}$		Determined with model (b)	
Mg^{2+} -PP1/2	Photolysis Ca^{2+} + Mg^{2+}	12	$2.5 \pm 0.6 \times 10^{-3}$	1.3×10^5	$3.0 \pm 0.9 \times 10^2$			
Ca^{2+} -NP-EGTA	Photolysis Ca^{2+}	4	$50 \pm 4 \times 10^{-9}$	$3.5 \pm 0.5 \times 10^7$	1.7 ± 0.2	$2.8 \pm 0.1 \times 10^{-6}$	35 ± 4	$10.3 \pm 0.7 \times 10^{-3}$
Ca^{2+} -PP	Photolysis Ca^{2+}	4	$10.7 \pm 0.1 \times 10^{-3}$	3.5×10^7	$3.6 \pm 0.1 \times 10^5$			

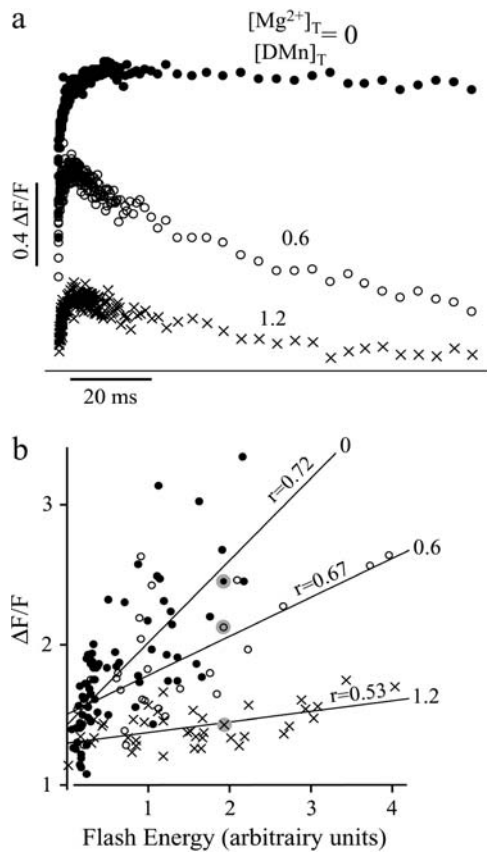


FIGURE 4 Fluorescence signals increase with increased UV flash energies. (a) Examples of uncaging with similar flash energies, but with different relative amounts of Mg^{2+} (○, no Mg^{2+} , i.e., $[\text{Mg}^{2+}]_{\text{T}}/[\text{DMn}]_{\text{T}} = 0$; ●, $[\text{Mg}^{2+}]_{\text{T}}/[\text{DMn}]_{\text{T}} = 0.6$; ×, $[\text{Mg}^{2+}]_{\text{T}}/[\text{DMn}]_{\text{T}} = 1.2$). (b) Linear fits to the plots of the maximum $\Delta F/F$ for different relative ratios of $[\text{Mg}^{2+}]_{\text{T}}/[\text{DMn}]_{\text{T}}$ as a function of flash energy. Symbols are the same as in a; the shaded dots indicate the maxima of the measurements shown in a.

and at comparable resting $[\text{Ca}^{2+}]$ (1–3 μM). The only difference between the traces is the total Mg^{2+} present. A rapid rise in $[\text{Ca}^{2+}]$ was observed, but in contrast to the Mg^{2+} -free experiments there was a strong drop in $[\text{Ca}^{2+}]$ after the initial rise (Fig. 4 a, solid versus open circles). In addition, at higher $[\text{Mg}^{2+}]$, larger flash energies were needed to produce commensurate changes in OGB-5N fluorescence after the flash, which is apparent from the decreased maximal fluorescence (Fig. 4, a and b). With increasing $[\text{Mg}^{2+}]$, more DMn will be occupied by Mg^{2+} and, therefore, less DMn will be available to bind Ca^{2+} and fewer Ca^{2+} ions will be needed to titrate the solution to the preflash $[\text{Ca}^{2+}]$ of 1–3 μM . This also means that fewer CaDMn molecules are present and less Ca^{2+} will be released for a certain flash energy. Furthermore, as Ca^{2+} is released, it will slowly replace some of the Mg^{2+} bound to DMn, hence the Ca^{2+} is buffered away after the uncaging.

To support the analysis of kinetic data in the presence of Mg^{2+} , we expanded the kinetic model to include interactions between DMn and Mg^{2+} (see Supplementary Material,

Appendix B). Every experiment started by uncaging Ca^{2+} in a solution without Mg^{2+} . From these data, the parameters for the interaction between DMn and Ca^{2+} were determined as described above for the given experiment. These parameters were then fixed at the averaged values obtained from three or four fits (each with a set of 10 traces). Different $[\text{Mg}^{2+}]_{\text{T}}$ were used that gave $[\text{Mg}^{2+}]_{\text{T}}/[\text{DMn}]_{\text{T}}$ ratios of 0.2, 0.4, 0.6, and 1.2. Using the $K_{\text{d}(\text{CaDMn})}$ and $K_{\text{d}(\text{MgDMn})}$ as found in the titration experiments, this would render an approximate resting $[\text{Mg}^{2+}]$ of ~0.1 mM, ~0.3 mM, ~0.5 mM, and ~1.5 mM, respectively. The different $[\text{Mg}^{2+}]_{\text{T}}/[\text{DMn}]_{\text{T}}$ ratios yielded comparable results and thus were grouped for statistical comparisons. We assumed that photolysis of Mg^{2+} -bound DMn (MgDMn) was similar to the photolysis of free DMn or of CaDMn (see the kinetic scheme shown in Fig. 5 a). This model will be referred to as “model (a)”. Fitting the Mg^{2+} data ($n = 12$ sets of traces, Fig. 6 a) with model (a) resulted in a $K_{\text{d}(\text{MgDMn})}$ of 0.5 ± 0.09 mM with an association rate ($k_{\text{on}(\text{MgDMn})}$) of $40 \pm 19 \times 10^3 \text{ M}^{-1} \text{ s}^{-1}$ and a dissociation rate ($k_{\text{off}(\text{MgDMn})}$) of $28 \pm 19 \text{ s}^{-1}$. For the photoproduct, the K_{d} for Mg^{2+} ($K_{\text{d}(\text{MgPP})}$) was $5.4 \pm 6.5 \text{ M}$, caused by a shift in the dissociation rate $k_{\text{off}(\text{MgDMn})}$ to a $k_{\text{off}(\text{MgPP})}$ of $92 \pm 53 \text{ ms}^{-1}$. Even though model (a) could reproduce the $\Delta F/F$ signals with Mg^{2+} present, the $K_{\text{d}(\text{MgDMn})}$ thus determined was ~300 times larger than both the $K_{\text{d}(\text{MgDMn})}$ found in our titration experiments ($p < 0.0001$) and the previously reported value for the dissociation constant of Mg^{2+} from DMn (Ayer and Zucker, 1999). If the $K_{\text{d}(\text{MgDMn})}$ was forced to a value within the range of its expected value (1–3 μM), no satisfactory fits could be obtained (not shown). Furthermore, as shown in the inset of Fig. 6 a, model (a) poorly reproduced the initial rise and fall in fluorescence. Thus, we conclude that model (a) does not accurately describe the actual uncaging process in these experiments.

Kaplan and Ellis-Davies (1988) showed that the PPs from DMn and CaDMn have different absorbance spectra. Therefore, the PPs resulting from either Ca^{2+} -bound or Ca^{2+} -free DMn are different, and the properties of these two PPs can differ significantly. In our experiments in the absence of Mg^{2+} , >99% of the DMn has bound Ca^{2+} and the contribution of the PP formed from free DMn is negligible. However, in experiments with Mg^{2+} present, a significant portion of DMn is occupied by Mg^{2+} . This MgDMn may also be uncaged into different PPs compared to the PPs formed by CaDMn. To test this hypothesis we evaluated the absorption spectra of PPs formed by DMn, CaDMn, and MgDMn. Fig. 5 c shows the absorbance spectra of DMn and the PPs formed from free DMn, CaDMn, and MgDMn. As expected (Kaplan and Ellis-Davies, 1988), the spectrum of DMn (Fig. 5 c, 1) was dramatically different from its PPs, and the spectra of the individual PPs were different from each other, depending on whether DMn was in its Ca^{2+} -free (Fig. 5 c, 2) or Ca^{2+} -bound (Fig. 5 c, 3) form before uncaging. More importantly, when Mg^{2+} was bound

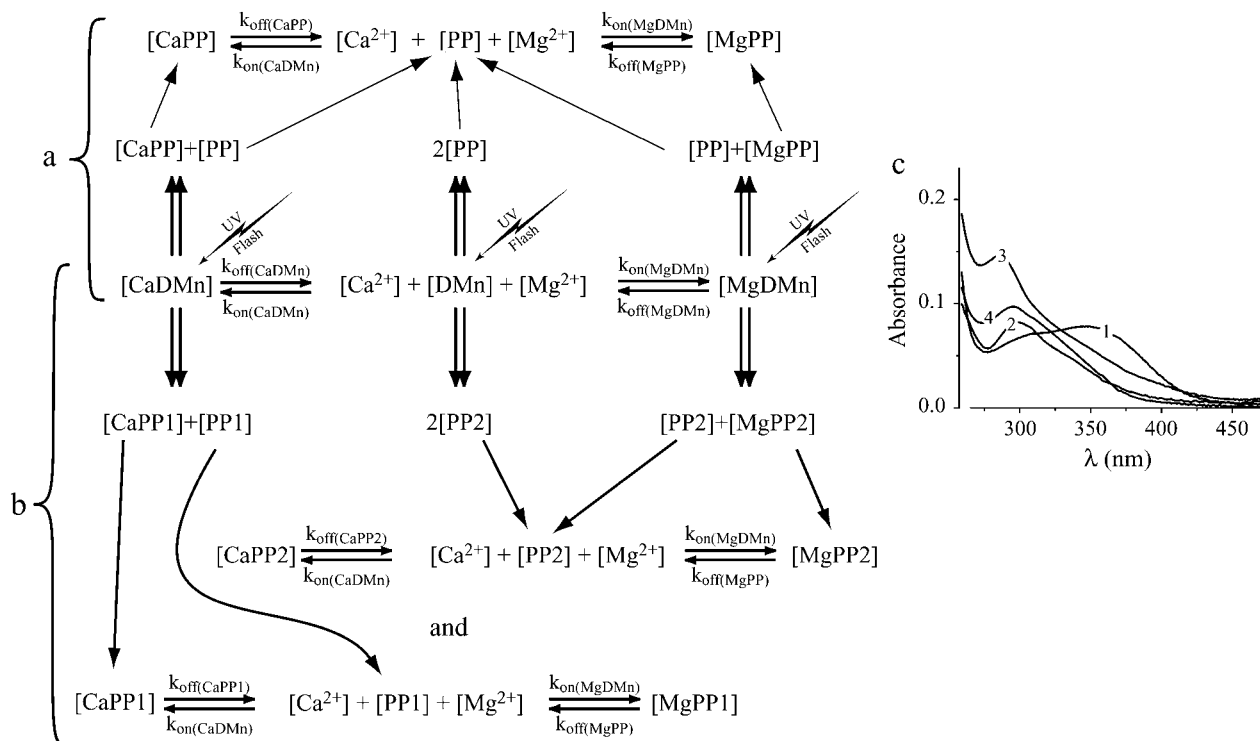


FIGURE 5 Schematics of the two models used to simulate Ca²⁺ and Mg²⁺ uncaging (double arrows) from DMn with the corresponding equilibria. (a) Model (a), in which DMn, CaDMn, and MgDMn all break up into similar photoproducts (PPs). (b) Model (b), in which CaDMn uncages into PP1, whereas DMn and MgDMn uncage into PP2 with a different Ca²⁺ affinity. (c) The difference in the absorption spectra for the PPs of DMn indicates that DMn forms different PPs depending on the ion it binds. The different numbers (1–4) indicate the following absorption spectra: 1, 25 μM DMn; 2, the PPs of 25 μM DMn; 3, the PPs of 25 μM CaDMn; and 4, the PPs of 25 μM MgDMn, all in high EDTA (>5 mM).

to DMn, the PPs had a significantly different spectra (Fig. 5 c, 4) compared to the PPs from CaDMn. This means that PPs formed from either CaDMn or MgDMn are different and may have different Ca²⁺ binding properties. Hence, we incorporated into our model the possibility that MgDMn uncages into different PPs than CaDMn (Fig. 5 b), and we will refer to this model as “model (b)”. In model (b), CaDMn breaks down into two PP1 molecules and MgDMn breaks down into two PP2 molecules. The Ca²⁺ binding properties of PP1 are determined from the experiments without Mg²⁺ present. The values for PP1 are fixed when analyzing experiments with Mg²⁺ present. It is nearly impossible to resolve if the PPs formed by free DMn can be represented by the properties of either PP1 or PP2. Furthermore, modeling the PPs of free DMn as PP1 or PP2 led to similar results, indicating a negligible contribution of free DMn. We chose to model the PPs of free DMn as PP2 because the absorption spectra of PPs of free DMn and MgDMn are alike. In addition, for simplicity we also assumed that the affinities and binding rate constants of either PP1 or PP2 for Mg²⁺ are similar. Fig. 5 b shows a schematic representation of this model. Fig. 6 b shows the same examples as Fig. 5 a fit with model (b), which more accurately describes the initial fast changes compared to model (a) (compare insets of Fig. 6, a and b). The $K_{d(\text{MgDMn})}$

determined with model (b) was $1.9 \pm 0.2 \mu\text{M}$ ($n = 12$), which changes to a $K_{d(\text{MgPP})}$ of $2.5 \pm 0.6 \text{ mM}$ (for both PP1 and PP2) after photolysis. For Ca²⁺, the $K_{d(\text{CaPP1})}$ was still 3.1 mM because it was determined without Mg²⁺ present, as described above. The PP2s formed from MgDMn had a $K_{d(\text{CaPP2})}$ of $89 \pm 41 \mu\text{M}$, a much higher affinity for Ca²⁺ than the PP1s formed by CaDMn. All association and dissociation rates are presented in Table 1 and are comparable to those found in studies using much longer illuminations of DMn (Ayer and Zucker, 1999) and also match the K_{d} s determined in our steady-state experiments. Apparently, DMn breaks up into different PPs depending on the ion bound to it. Earlier findings (Ayer and Zucker, 1999; Kaplan and Ellis-Davies, 1988) and our own results suggest that at least two types of PPs are formed, one with a K_d for Ca²⁺ of $\sim 3 \text{ mM}$ and another with a K_d of $\sim 10\text{--}100 \mu\text{M}$.

DISCUSSION

We studied the properties of NP-EGTA, a Ca²⁺ cage with negligible Mg²⁺ affinity, to determine its usefulness in inducing step-like changes in [Ca²⁺] under our desired experimental conditions (resting [Ca²⁺], 1–3 μM; resting [Mg²⁺], 1 mM). We also determined the Ca²⁺ and Mg²⁺ binding kinetics to DMn and its PPs using ultrafast flash

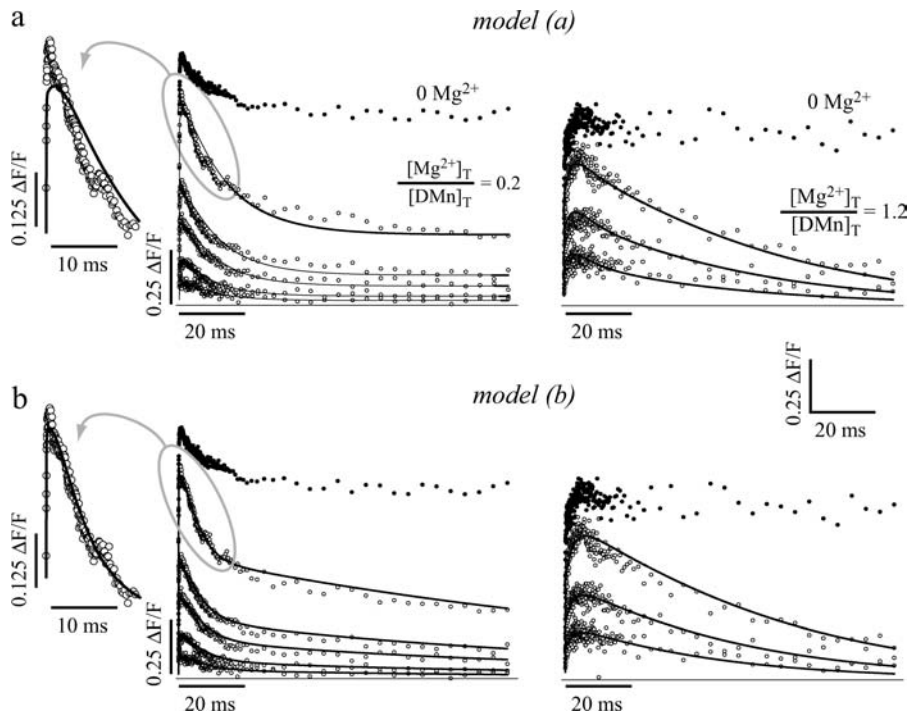


FIGURE 6 Model-generated fits of Ca^{2+} uncaging from DMn with Mg^{2+} present. Examples of data obtained when uncaging Ca^{2+} in the presence of Mg^{2+} (\circ). Left panel, $[\text{Mg}^{2+}]_{\text{T}}/[\text{DMn}]_{\text{T}} = 0.2$; right panel, $[\text{Mg}^{2+}]_{\text{T}}/[\text{DMn}]_{\text{T}} = 1.2$. For comparison, the top traces in each data set represent Ca^{2+} uncaging in the absence of Mg^{2+} with a comparable $\Delta F/F$ value (\bullet ; fitted curves have been omitted for clarity). (a) Fits using model (a). (b) Fits using model (b). Insets show on a faster timescale the largest curves with Mg^{2+} present. It is clear that model (b) is far better than model (a) in fitting the fast initial rise and fall in fluorescence.

photolysis. Once the Ca^{2+} binding properties of DMn were determined in the absence of Mg^{2+} , the kinetic parameters were used in an extended model incorporating Mg^{2+} binding to DMn. The results of our measurements will enable the use of Ca^{2+} uncaging experiments in the presence of physiological $[\text{Mg}^{2+}]$ to resolve the properties of Ca^{2+} binding kinetics to various CBPs under physiological conditions.

Uncaging time constants of DMn and NP-EGTA

Our model accurately described the uncaging of DMn and NP-EGTA and their Ca^{2+} binding if only two uncaging time constants were used. Earlier studies of DMn, NP-EGTA, or similar compounds irradiated by short UV pulses have shown that the resulting *aci*-nitro intermediates (intermediate molecules formed in the transition from DMn or NP-EGTA to PPs) have two decay rates (Ellis-Davies et al., 1996; Il'ichev et al., 2004). For example, for DMn at pH 7.2, 66% of the *aci*-nitro species decays at a rate of $80,000 \text{ s}^{-1}$, whereas the rest decays at $11,000 \text{ s}^{-1}$. However, since the rate-limiting step of the entire uncaging process is unknown, and may vary depending on pH (Il'ichev et al., 2004; Schupp et al., 1987), the decay rates of any given intermediate *aci*-nitro species may not reflect the rate of the complete uncaging process. Nevertheless, the two different pathways for uncaging DMn (Ellis-Davies et al., 1996) could underlie the two uncaging time constants found in our experiments. The ratios of the $\tau_{\text{fast}}/\tau_{\text{slow}}$ components of 67:33% for DMn and 36:64% for NP-EGTA are in good agreement with the 66:34% and 37:63% ratios found for the distributions of

decay rates for the *aci*-nitro species for DMn and NP-EGTA, respectively (Ellis-Davies et al., 1996).

Changes in pH during uncaging may be an alternative explanation for two uncaging time constants in our experiments. Proton binding in the last transition step from *aci*-nitro species to PPs can raise local pH to modulate the pH-sensitive uncaging process (Ellis-Davies et al., 1996; Il'ichev et al., 2004). This latter explanation seems unlikely since there is an excess pH buffer present (40 mM HEPES) and only small (local) amounts of DMn are uncaged (<1.5%).

Previously recorded Ca^{2+} signals after DMn uncaging were adequately described with models using only one uncaging time constant (Ayer and Zucker, 1999; Ellis-Davies et al., 1996; Escobar et al., 1997), but under quite different experimental conditions. For example, Ellis-Davies et al. (1996) used fast recordings at high resting $[\text{Ca}^{2+}]$ ($\sim 70 \mu\text{M}$) to determine the uncaging time constants of DMn and NP-EGTA. The short timescale ($\sim 200 \mu\text{s}$) of their measurements precludes distinguishing a second slower time constant. In their experiments on a longer timescale ($\sim 3 \text{ ms}$), the resting $[\text{Ca}^{2+}]$ used ($\sim 15 \text{ nM}$ for DMn and $\sim 450 \text{ nM}$ for NP-EGTA) was much lower than in ours, which was also the case in earlier experiments by Escobar et al. (1997). Under such conditions, considerable amounts of DMn or NP-EGTA are Ca^{2+} -free and can mask slow Ca^{2+} release by acting as relatively fast Ca^{2+} buffers. We tested this hypothesis by simulating experiments using our model (with two uncaging time constants) under conditions equivalent to those used by Ellis-Davies et al. (1996). In these simulations the data could be fitted with single uncaging time constants

for DMn or NP-EGTA similar to those found by Ellis-Davies et al. (1996). Thus, a model with only one uncaging time constant might suffice depending on the experimental conditions. However, other conditions such as ours will require a second uncaging time constant to accurately describe DMn or NP-EGTA uncaging.

DMn and Ca^{2+}

The $K_{d(\text{CaDMn})}$ found by our fitting method (7.6 nM) is in good agreement with the steady state experiments either with or without Mg^{2+} present (7.0 nM and 8.1 nM, respectively) and is comparable to previously measured or estimated values of ~ 5 nM (Escobar et al., 1997; Kaplan and Ellis-Davies, 1988). Compared to the three reported on-rates for Ca^{2+} binding to DMn— $\sim 8 \times 10^7 \text{ M}^{-1} \text{ s}^{-1}$ (Ellis-Davies et al., 1996), $\sim 3 \times 10^7 \text{ M}^{-1} \text{ s}^{-1}$ (Ayer and Zucker, 1999; Escobar et al., 1997; Xu et al., 1997), and $\sim 1.5 \times 10^6 \text{ M}^{-1} \text{ s}^{-1}$ (Zucker, 1993)—the $k_{\text{on}(\text{CaDMn})}$ of $2.9 \times 10^7 \text{ M}^{-1} \text{ s}^{-1}$ found in our experiments strongly supports intermediate binding rates. For the PPs we found that the $K_{d(\text{CaPP})}$ was 3.1 mM, again in close agreement with the initially reported and widely used value of ~ 3 mM (Ellis-Davies et al., 1996; Escobar et al., 1995, 1997; Kaplan and Ellis-Davies, 1988; Xu et al., 1997). However, Neher and Zucker (1993) reported a sixfold lower K_d of 250 μM . It was later suggested that this could be an apparent K_d resulting from a combination of the two dissimilar K_d s of the two PPs (PP1 and PP2) formed from CaDMn and DMn (Ayer and Zucker, 1999).

DMn and Mg^{2+}

The $K_{d(\text{MgDMn})}$ of 1.5 μM determined in our titration experiments is comparable to the originally estimated 2.5 μM (Kaplan and Ellis-Davies, 1988). Our model (a) resulted in a much higher estimate of $K_{d(\text{MgDMn})}$ (0.5 mM) and this model also failed to fit the initial phase of the Ca^{2+} signals. We incorporated a second PP, formed by MgDMn and free DMn, with a higher affinity for Ca^{2+} but otherwise similar properties for Mg^{2+} , thus creating model (b), which accurately fitted our results. Based on the absorbance spectra we show that the PPs formed from MgDMn are indeed different from those formed by CaDMn.

Our findings are in many respects similar to those of Ayer and Zucker (1999): PPs formed by MgDMn have a much higher affinity for Ca^{2+} than PPs formed by CaDMn. We found the $K_{d(\text{CaPP2})}$ to be $\sim 90 \mu\text{M}$, within the same order of magnitude as the $\sim 13 \mu\text{M}$ estimated by Ayer and Zucker (1999). Also in agreement with Ayer and Zucker (1999), the association rate of Mg^{2+} binding is slower than that for Ca^{2+} , since Mg^{2+} binding is rate-limited by the removal of H_2O from hydrated Mg^{2+} (Hague, 1977). The most significant difference between the previous model (Ayer and Zucker, 1999) and ours for the uncaging of DMn or NP-

EGTA is that we use two uncaging time constants instead of one.

Assumptions

In a prior analysis (Ayer and Zucker, 1999), various constants, including $K_{d(\text{CaDMn})}$, $k_{\text{on}(\text{CaDMn})}$, $K_{d(\text{CaPP1})}$, and $K_{d(\text{MgPP})}$, were preset to previously reported values. Our approach determined the most critical values by iteration with the model, testing its robustness. Nevertheless, like others (Ayer and Zucker, 1999; Ellis-Davies et al., 1996; Escobar et al., 1995, 1997; Kaplan and Ellis-Davies, 1988; Nagerl et al., 2000), we had to make some assumptions about how DMn uncages. One of the major assumptions is that every DMn molecule breaks up into two PPs with similar properties. Based on the general idea of how DMn breaks up according to Fig. 2 a (adapted from Kaplan and Ellis-Davies, 1988), the two fractions should be very different. If DMn can uncage in different ways into PP1 and PP2, there should be up to four different photoproducts, possibly with different properties. Based on our data, it is impossible to resolve these reactions in such detail, but our model has proven to be sufficiently accurate to be applicable to the experimental data. Previously, we and others used a simpler model in experiments with only Ca^{2+} present, where every DMn molecule “broke up” into only one molecule of photoproduct with a significant affinity for Ca^{2+} (Escobar et al., 1995, 1997; Nagerl et al., 2000). We compared this earlier model to our expanded models used here. Surprisingly, the results for fitting with the simpler model were quite similar to the results obtained here. Since only small amounts of the total DMn ($\sim 1\%$) are uncaged in our experiments, the relative amounts of the resulting PPs will be nearly equal.

Another assumption made by us and others (Ayer and Zucker, 1999; Ellis-Davies et al., 1996; Nagerl et al., 2000) is that the change in affinity between DMn or NP-EGTA and its PPs is caused by a change in the off-rates, whereas the on-rates remain constant. It was also assumed that the quantum yield and kinetics for uncaging Ca^{2+} -bound, Mg^{2+} -bound, or ion-free DMn or NP-EGTA are equal. Apparently, small errors possibly introduced by these latter assumptions do not significantly interfere with the accuracy of the model.

DMn versus NP-EGTA

Our goal was to induce steplike changes in $[\text{Ca}^{2+}]$ by uncaging from DMn in the presence of Mg^{2+} . An accurate model for the underlying uncaging kinetics will be essential to determine Ca^{2+} binding kinetics to CPBs in the presence of Mg^{2+} . But is the use of DMn warranted when other Ca^{2+} cages with no sensitivity to Mg^{2+} (e.g., NP-EGTA and nitr-5) are commercially available? The relatively high affinities for Ca^{2+} (6.3 μM (Ellis-Davies, 2003)) of the photoproducts of nitr-5 make this Ca^{2+} cage of limited use for inducing steplike changes in $[\text{Ca}^{2+}]$. However, the properties of

NP-EGTA might render it useful for Ca^{2+} photolysis in the presence of Mg^{2+} . Using the determined properties of DMn and NP-EGTA, we simulated the change in $[\text{Ca}^{2+}]$ after uncaging from either of these compounds under different experimental conditions (Fig. 7). Despite the relatively high affinity of DMn for Mg^{2+} , we found significant advantages of using DMn over NP-EGTA under some conditions.

When studying Ca^{2+} binding to CBPs (Nagerl et al., 2000), the experiments have to be performed at a relatively high resting $[\text{Ca}^{2+}]$ to minimize rebinding of uncaged Ca^{2+} to free DMn or NP-EGTA. Fig. 7, *a* and *b*) shows simulations of Ca^{2+} uncaging from 5 mM DMn or NP-EGTA solutions at a resting $[\text{Ca}^{2+}]$ of 1 μM to induce steplike changes in $[\text{Ca}^{2+}]$. With no Mg^{2+} present (Fig. 7 *a*), 1% uncaging of DMn will produce an $\sim 18\text{-}\mu\text{M}$ step in $[\text{Ca}^{2+}]$ (Fig. 7 *a* and *b*, *I*). More uncaging is needed to evoke similar long lasting changes in $[\text{Ca}^{2+}]$ from NP-EGTA

(compare Fig. 7 *a*, *II* and *III*). However, the major disadvantage of NP-EGTA is that only $\sim 35\%$ of it uncages rapidly ($\tau = \sim 3 \mu\text{s}$) and $\sim 65\%$ uncages slowly ($\tau = \sim 10 \text{ ms}$). Consequently, no steplike change in $[\text{Ca}^{2+}]$ can be obtained. Instead, a biphasic change in $[\text{Ca}^{2+}]$ results that is more evident with more uncaging (Fig. 7 *a*, *IV* and *V*). This means that under these conditions Ca^{2+} uncaging from NP-EGTA makes it impossible to study Ca^{2+} binding kinetics of various processes that respond to higher $[\text{Ca}^{2+}]$, or of those that temporally overlap with the dip in the $[\text{Ca}^{2+}]$ caused by Ca^{2+} rebinding to NP-EGTA. Therefore, the use of NP-EGTA would be limited to the study of processes faster than the rebinding of Ca^{2+} to NP-EGTA ($k_{\text{on}} = 3.5 \times 10^5 \text{ M}^{-1} \text{ s}^{-1}$) or to those slower than the slow time constant of uncaging ($\tau = 10 \text{ ms}$). This problem may be avoided by using higher resting $[\text{Ca}^{2+}]$ to lower the free NP-EGTA available to rebind the released Ca^{2+} . An $\sim 99.5\%$ occupancy

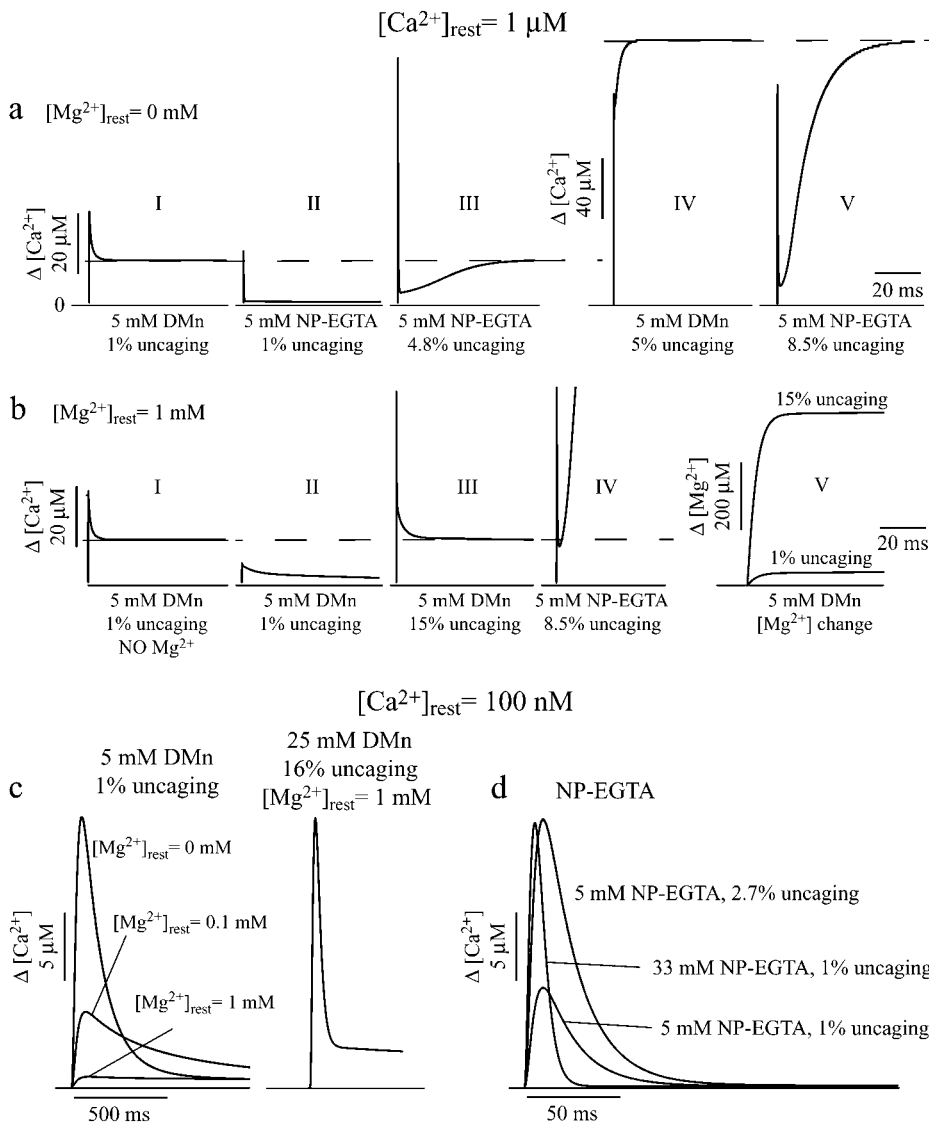


FIGURE 7 Simulations of Ca^{2+} uncaging using model (b) and the parameters obtained from our measurements (also see Table 1). (a) Uncaging of 5 mM DMn or NP-EGTA with an initial $[\text{Ca}^{2+}]$ of 1 μM and no Mg^{2+} present. To reach a similar steady-state $[\text{Ca}^{2+}]$ level, more NP-EGTA has to be uncaged. Release of Ca^{2+} from NP-EGTA release causes a clear “dip” in $[\text{Ca}^{2+}]$. (b) At a resting $[\text{Mg}^{2+}]$ of 1 mM, Ca^{2+} release is significantly reduced (compare panels *I* and *II*), but can be restored by increasing the fraction of uncaging (panel *III*). Uncaging from NP-EGTA to produce a similar initial rise in $[\text{Ca}^{2+}]$ will result in a large secondary rise in $[\text{Ca}^{2+}]$ (panel *IV*). With Mg^{2+} present, uncaging from DMn will also release considerable amounts of Mg^{2+} as shown in panel *V*. (c) At the physiological resting $[\text{Ca}^{2+}]$ of 100 nM, the amount of uncaged Ca^{2+} from DMn decreases significantly when Mg^{2+} is added; however, peak values of the $[\text{Ca}^{2+}]$ spike can be restored by increasing the level of uncaging (*c*, right panel). (d) Uncaging from NP-EGTA can accomplish a similar peak in $[\text{Ca}^{2+}]$ if the uncaging level is slightly increased or the total NP-EGTA is considerably elevated. However, note that the Ca^{2+} spikes after release from NP-EGTA are much shorter lasting than those after uncaging DMn (note the 10-fold difference between the timescales of *c* and *d*).

of DMn is sufficient to eliminate any major drop in $[Ca^{2+}]$ directly after uncaging (Figs. 2, b and e, and 5). A resting free $[Ca^{2+}]$ of $\sim 10 \mu M$ would be needed to accomplish 99.5% occupancy of NP-EGTA. Such a high concentration may confound the study of a process with a high-affinity Ca^{2+} component. For instance, CBPs with a K_d of 525 nM or less (Nagerl et al., 2000), will have $>95\%$ of these sites occupied at $10 \mu M [Ca^{2+}]$. Thus it might be difficult to detect changes in $[Ca^{2+}]$ produced by $<5\%$ of the free binding sites. In comparison, at $1.4 \mu M$ free $[Ca^{2+}]$ (99.5% occupancy of DMn) the 27% unoccupied binding sites of a CBP with a K_d of ~ 500 nM would more likely produce detectable Ca^{2+} changes.

The use of DMn is advantageous even when Mg^{2+} is present. Large uncaging energies are needed, but small step-like changes in $[Ca^{2+}]$ can still be made (compare Fig. 7 b, I–III). It may be argued that NP-EGTA may be used for experiments requiring small changes in $[Ca^{2+}]$, as long as the dip in $[Ca^{2+}]$ is above the minimal desired value of $[Ca^{2+}]$. However, the $[Ca^{2+}]$ would still steadily rise after this dip (Fig. 7 b, IV), complicating the measurements. It should be taken into account that uncaging DMn with Mg^{2+} present will also release considerable amounts of Mg^{2+} (Fig. 7 b, V), most of it directly by uncaging MgDMn and a very small amount (estimated at $<1\%$) by uncaged Ca^{2+} ions replacing Mg^{2+} ions bound to nonphotolyzed DMn molecules. This also induces a slow decay in $[Ca^{2+}]$ (Fig. 5), but this drop in $[Ca^{2+}]$ is too slow to interfere with the faster binding processes expected to be present in CPBs (Nagerl et al., 2000).

To replicate possible physiological uncaging experiments, we also simulated Ca^{2+} uncaging at the intracellular physiological resting $[Ca^{2+}]$ of 100 nM (Fig. 7, c and d), with 5 mM total DMn and 1% uncaging. As previously reported (Escobar et al., 1997; Zucker, 1993), with no Mg^{2+} present, uncaging results in a sharp and short-lived Ca^{2+} spike of $\sim 20 \mu M$ and within ~ 0.7 ms the $[Ca^{2+}]$ reaches a steady-state value of ~ 240 nM (Fig. 7 a). Upon addition of only 0.1 mM free $[Mg^{2+}]$, the $[Ca^{2+}]$ peak becomes dramatically reduced to $\sim 6 \mu M$ and declines within ~ 2 ms to a steady-state $[Ca^{2+}]$ of ~ 310 nM. Increasing the free $[Mg^{2+}]$ to 1 mM, the high end of its intracellular concentration (Romani and Scarpa, 2000, Li-Smerin et al., 2001) reduces the peak and steady-state levels to ~ 850 nM and ~ 460 nM, respectively. The simulations indicate that the peak $[Ca^{2+}]$ is inversely related to the relative amount of DMn occupied by Mg^{2+} (not shown). To compensate for the diminished $[Ca^{2+}]$ change without increasing the resting $[Ca^{2+}]$, either the uncaging percentage (flash energy) or the DMn concentration should be increased. With 1 mM $[Mg^{2+}]$ changing just one of these parameters would not be sufficient to evoke an $\sim 6\text{-}\mu M$ peak in $[Ca^{2+}]$, unless a concentration of >500 mM DMn is used or an uncaging of 80% is achieved. The right panel in Fig. 7 c shows Ca^{2+} release from 16% uncaging with 25 mM DMn starting at a $[Mg^{2+}]$ of 1 mM

and a $[Ca^{2+}]$ of 100 nM. A combined increase in total DMn and uncaging can rescue the peak $[Ca^{2+}]$, although the peak will become shorter and the steady-state $[Ca^{2+}]$ will be higher. With 5 mM NP-EGTA and 100 nM resting $[Ca^{2+}]$, 1% uncaging obviously does not create a Ca^{2+} pulse with the same amplitude as 1% DMn uncaging. The peak of the Ca^{2+} pulse can be increased to $\sim 20 \mu M$ by augmenting the release to $\sim 2.7\%$ or by increasing the NP-EGTA concentration to ~ 33 mM. Therefore, NP-EGTA seems to be a good candidate to induce intracellular Ca^{2+} spikes with Mg^{2+} present. However, compared to Ca^{2+} release from DMn, the Ca^{2+} spikes are very short-lasting. The Ca^{2+} spike created by 2.7% uncaging NP-EGTA lasts $\sim 50 \mu s$ (Fig. 7, c and d) compared to $500 \mu s$ for 1% uncaging from DMn with no Mg^{2+} present. Various Ca^{2+} -dependent processes may not sense such extremely short Ca^{2+} pulses because of their relatively slow association rates. In conclusion, it may be problematic to use DMn in environments with low resting $[Ca^{2+}]$ and physiological $[Mg^{2+}]$, but NP-EGTA may not be a perfect alternative, as it produces extremely short-lived changes in $[Ca^{2+}]$. The simulations with DMn uncaging with Mg^{2+} present indicate that Mg^{2+} will also be released (not shown). The $[Mg^{2+}]$ rises slowly after the flash, but it is relatively small: with 1 mM Mg^{2+} and 1% uncaging the $[Mg^{2+}]$ would rise by $45 \mu M$ after 100 ms.

CONCLUSIONS

Our model was used to predict the behavior of DMn near physiological levels of free $[Ca^{2+}]$ (100 nM) and $[Mg^{2+}]$ (1 mM). Its main limitation as a Ca^{2+} cage is not the excessive release of Mg^{2+} , but the small $[Ca^{2+}]$ rise produced under these conditions. If the goal is to produce relatively small changes in cellular $[Ca^{2+}]$, DMn is suitable, particularly since only negligible changes in Mg^{2+} will occur. Intracellular Ca^{2+} uncaging from NP-EGTA might be a better solution, if very short-lived Ca^{2+} signals are desirable. In other experiments, such as determination of binding rates of CPBs, where a higher initial $[Ca^{2+}]$ might be tolerated or even preferable, step-like changes in $[Ca^{2+}]$ can best be produced by DMn even in the presence of Mg^{2+} .

SUPPLEMENTARY MATERIAL

An online supplement to this article can be found by visiting BJ Online at <http://www.biophysj.org>.

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Appendix A

For the steady state equilibrium we considered equation 1 to express $[CaDMn]/ [DMn]_T$ as a function of $[Ca^{2+}]$, $[Ca^{2+}]_T$, $[Mg^{2+}]_T$, $[DMn]_T$, $K_{d(CaDMn)}$ and, $K_{d(MgDMn)}$.

The fraction of Ca^{2+} bound DMn is:

$$\frac{[CaDMn]}{[DMn]_T} = \frac{[CaDMn]}{[CaDMn] + [MgDMn] + [DMn]} \quad (1)$$

where

$$[CaDMn] = \frac{[Ca^{2+}] \cdot [DMn]}{K_{d(CaDMn)}} \quad (2)$$

$$[MgDMn] = \frac{[Mg^{2+}] \cdot [DMn]}{K_{d(MgDMn)}} \quad (3)$$

which gives:

$$\frac{[CaDMn]}{[DMn]_T} = \frac{K_{d(MgDMn)} \cdot [Ca^{2+}]}{K_{d(MgDMn)} \cdot [Ca^{2+}] + K_{d(CaDMn)} \cdot [Mg^{2+}] + K_{d(CaDMn)} \cdot K_{d(MgDMn)}} \quad (4)$$

Similar to equation 6 we derive

$$\frac{[MgDMn]}{[DMn]_T} = \frac{K_{d(CaDMn)} \cdot [Mg^{2+}]}{K_{d(MgDMn)} \cdot [Ca^{2+}] + K_{d(CaDMn)} \cdot [Mg^{2+}] + K_{d(CaDMn)} \cdot K_{d(MgDMn)}} \quad (5)$$

$$[MgDMn] = [Mg^{2+}]_T - [Mg^{2+}] \quad (6)$$

Equations 7 and 8 can be rewritten as (for this, we used MatLab student version 5.3):

$$[Mg^{2+}] = \frac{K_{d(CaDMn)} \cdot [Mg^{2+}]_T - K_{d(MgDMn)} (K_{d(CaDMn)} - [Ca^{2+}] - [DMn]_T) + X}{2K_{d(CaDMn)}} \quad (7a)$$

where

$$X = \sqrt{K_{d(CaDMn)}^2 ([Mg^{2+}]_T ([Mg^{2+}]_T + 2 \cdot K_{d(MgDMn)} - 2 \cdot [DMn]_T) + [DMn]_T ([DMn]_T + 2 \cdot K_{d(MgDMn)}) + K_{d(MgDMn)}^2) + K_{d(MgDMn)} \cdot [Ca^{2+}] (K_{d(MgDMn)} \cdot [Ca^{2+}] + 2 \cdot K_{d(CaDMn)} ([Mg^{2+}]_T + K_{d(MgDMn)} [DMn]_T))} \quad (9b)$$

Substitution of $[Mg^{2+}]$ in eq. 6 with eq. 9 gives $[CaDMn]/[DMn]_T$ as a function of the measured parameter $[Ca^{2+}]$, and the known parameters $[Ca^{2+}]_T$, $[Mg^{2+}]_T$ and, $[DMn]_T$ and the unknown parameters $K_{d(CaDMn)}$ and $K_{d(MgDMn)}$.

Appendix B

The Berkeley Madonna model:

$[Ca^{2+}]_T, [Ca^{2+}]_{t=0}$	total Ca^{2+} concentration (calc.), free Ca^{2+} concentration (sim.) and, initial free Ca^{2+} concentration (measured with Ca^{2+} -sensitive electrode), respectively
$[Mg^{2+}]_T, [Mg^{2+}]_{t=0}$	total Mg^{2+} concentration (known), free Mg^{2+} concentration (sim.) and, initial free Mg^{2+} concentration (calc.), respectively
$[DMn]_T, [DMn]_{t=0}$	total DMn concentration (measured with photospectrometer), free DMn concentration (sim.) and, initial DMn concentration (calc.), respectively
$K_{d(CaDMn)}$	equilibrium dissociation constant of DMn with Ca^{2+} (fitted in set of initial experiments*)
$k_{on(CaDMn)}$	association rate constant of DMn with Ca^{2+} (calculated = $k_{off(CaDMn)} / K_{d(CaDMn)}$)
$k_{off(CaDMn)}$	dissociation rate constant of DMn with Ca^{2+} (fitted in set of initial experiments*)
$K_{d(MgDMn)}$	equilibrium dissociation constant of DMn with Mg^{2+} (fitted)
$k_{on(MgDMn)}$	association rate constant of DMn with Mg^{2+} (calculated = $k_{off(MgDMn)} / K_{d(MgDMn)}$)
$k_{off(MgDMn)}$	dissociation rate constant of DMn with Mg^{2+} (fitted)
$[CaDMn], [CaDMn]_{t=0}$	concentration of Ca^{2+} -DMn complex (sim.), and initial concentration (calc), resp.
$[MgDMn], [MgDMn]_{t=0}$	concentration of Mg^{2+} -DMn complex (sim.), and initial concentration (calc), resp.
$[PP1], [PP2]$	photoproduct concentration either formed by uncaging of CaDMn (PP1) or DMn or MgDMn (PP2), initial values are 0, (sim.)
$K_{d(CaPP1)}, K_{d(CaPP2)}$	equilibrium dissociation constant of PP1 and PP2, respectively, with Ca^{2+} (PP1 is fitted in set of initial experiments*, PP2 is fitted later)
$k_{off(CaPP1)}, k_{off(CaPP2)}$	equilibrium dissociation constant of PP1 and PP2, respectively, with Ca^{2+} (PP1 is fitted in set of initial experiments*, PP2 is fitted later)
$K_{d(MgPP)}$	equilibrium dissociation constant of both PP1 and PP2 with Mg^{2+} (fitted)
$k_{off(MgPP)}$	equilibrium dissociation constant of both PP1 and PP2 with Mg^{2+} (fitted)
$[CaPP1], [CaPP2]$	concentration of Ca^{2+} -PP1 complex (sim.) and Ca^{2+} -PP2 complex (sim.), resp.
$[MgPP1], [MgPP2]$	concentration of Mg^{2+} -PP1 complex (sim.) and Mg^{2+} -PP2 complex (sim.), resp.
α, x	fraction of DM-n to be photolysed (fitted per trace), and its fraction that photolyzes with the fast time constant (fitted in set of initial experiments*).
$[DMn]_f, [CaDMn]_f, [MgDMn]_f$	total DMn concentration., total CaDMn concentration and, total MgDMn concentrations that photolyzes with the fast timeconstant (all calc.), respectively
$[DMn]_s, [CaDMn]_s, [MgDMn]_s$	total DMn concentration., total CaDMn concentration and, total MgDMn concentrations that photolyzes with the slow timeconstant (all calc.), respectively
τ_{fast}	time constant of fast photolysis (fitted in set of initial experiments*)
τ_{slow}	time constant of slow photolysis (fitted in set of initial experiments*)
$J_{fast(DMn)}, J_{fast(CaDMn)}, J_{fast(MgDMn)}$	flux/rate of fast uncaging to the various photoproducts and complexes of DMn, CaDMn and, MgDMn, respectively (all simulated)

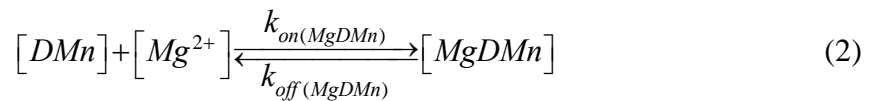
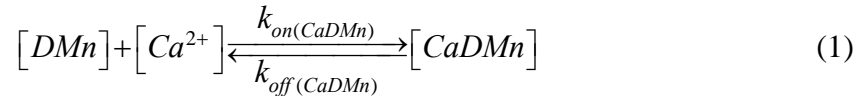
$J_{slow(DMn)}, J_{slow(CaDMn)}, J_{slow(MgDMn)}$	flux/rate of slow uncaging to the various photoproducts and complexes of DMn, CaDMn and, MgDMn, respectively (all simulated)
t_{flash}	moment of initiation of photolysis (in these simulations 0)
$[D]_T$	total dye (Ca^{2+} -indicator) concentration = 100 μ M OGB-5N
$[D]$	concentration of free dye (sim.)
$[CaD]$	concentration of Ca^{2+} -dye complex (sim.)
$K_{d(D)}$	equilibrium dissociation constant of the dye, OGB-5N = 29.3 μ M (measured with Ca^{2+} solutions of known $[Ca^{2+}]$)
$k_{on(D)}$	association rate constant of the dye, OGB-5N with $Ca^{2+} = 2.6 \cdot 10^8 \text{ M}^{-1} \text{ s}^{-1}$ (calculated = $k_{off(D)} / K_{d(D)}$)
$k_{off(D)}$	dissociation rate constant of the dye, OGB-5N with $Ca^{2+} = 7.5 \times 10^3 \text{ s}^{-1}$ or $8.7 \times 10^3 \text{ s}^{-1}$ depending on the lot used (measured, derived from relaxation time constant of Ca^{2+} -dye complex after Ca^{2+} -pulse)
F_{ratio}	ratio F_{max} / F_{min} for the dye, OGB-5N = 10.8 or 40.0 depending on the lot used (measured with Ca^{2+} solutions of known $[Ca^{2+}]$)

DMn can be replaced everywhere for NP-EGTA to represent the uncaging model for NP-EGTA

The scheme of uncaging Ca^{2+} and Mg^{2+} and detecting Ca^{2+}

Pre flash

The kinetic reactions for the different fractions before photolysis are:



Flash photolysis of DM-nitrophen

Of the total DM-nitrophen (DMn) concentration ($[DMn]_T = [DMn], [CaDMn], [MgDMn]$) a certain fraction α photolyzes. The flash energy of the UV laser determines the total proportion (α). Of this fraction α a fraction (x) photolyzes rapidly ($[DM]_f, [CaDMn]_f, [MgDMn]_f$) while a fraction ($1-x$) photolyzes slowly ($[DMn]_s, [CaDMn]_s, [MgDMn]_s$)

$$[DMn]_f = \alpha x [DMn] \quad (3)$$

$$[CaDMn]_f = \alpha x [CaDMn] \quad (4)$$

$$[MgDMn]_f = \alpha x [MgDMn] \quad (5)$$

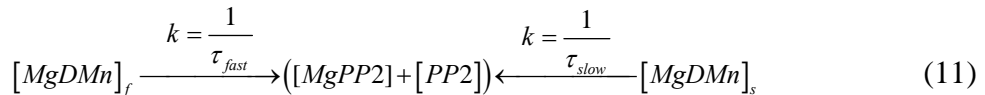
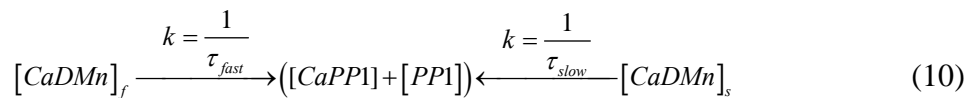
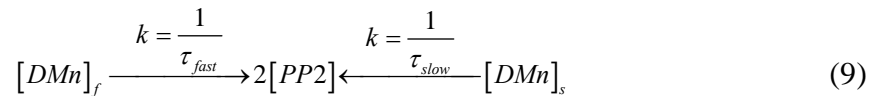
$$[DMn]_s = \alpha (1-x) [DMn] \quad (6)$$

$$[CaDMn]_s = \alpha (1-x) [CaDMn] \quad (7)$$

$$[MgDMn]_s = \alpha (1-x) [MgDMn] \quad (8)$$

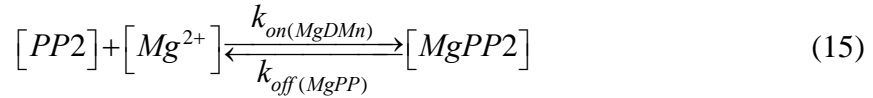
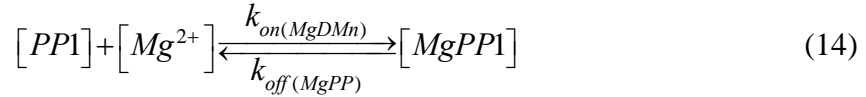
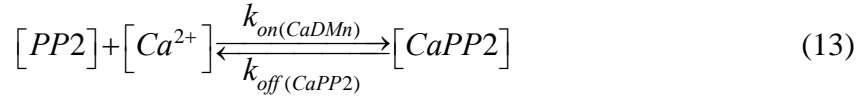
Where $x=1$ for one uncaging timeconstant and $x \leq 1$ for two uncaging timeconstants.

After the flash the affected DMn molecules transition into photoproducts of DMn ($[PP2]$), CaDMn ($[CaPP1]$ and $[PP1]$) and, MgDMn ($[MgPP2]$ and $[PP2]$) with time constants τ_{fast} and τ_{slow} :



In eq. 18 as Ca²⁺-free DMn is broken down into PP2 as described earlier by {Ayer, 1999 192 /id;Kaplan, 1988 179 /id}. However, based on our experiments it also be broken down into PP1.

Compared to DMn, the individual photoproducts may have different dissociation rate constants for Ca²⁺ and Mg²⁺ (leading to alternate affinities):



where we can note that the dissociation rates for PP1 and PP2 to Mg^{2+} are similar ($k_{offMgPP}$)

whereas for Ca^{2+} they depend on the model used:

$$\text{model (a): } k_{offCaPP1} = k_{offCaPP2} \quad (16)$$

$$\text{model (b): } k_{offCaPP1} \neq k_{offCaPP2} \quad (17)$$

Determining $[Ca^{2+}]$ with a Ca^{2+} -indicator

The kinetic reaction for the fluorescent dye (D) is:



In the presence of both DMn and the dye there is a complete model for uncaging Ca^{2+} and Mg^{2+} and detecting Ca^{2+} .

Equations for uncaging Ca^{2+} and Mg^{2+} and detecting Ca^{2+}

For DMn

$$\frac{d[\text{CaDMn}]}{dt} = k_{on(\text{CaDMn})} \cdot [\text{Ca}^{2+}] \cdot [\text{DMn}] - k_{off(\text{CaDMn})} \cdot [\text{CaDMn}] - J_{fast(\text{CaDMn})} - J_{slow(\text{CaDMn})} \quad (19)$$

$$\frac{d[\text{MgDMn}]}{dt} = k_{on(\text{MgDMn})} \cdot [\text{Mg}^{2+}] \cdot [\text{DMn}] - k_{off(\text{MgDMn})} \cdot [\text{MgDMn}] - J_{fast(\text{MgDMn})} - J_{slow(\text{MgDMn})} \quad (20)$$

$$\begin{aligned} \frac{d[\text{DMn}]}{dt} = & -k_{on(\text{CaDMn})} \cdot [\text{Ca}^{2+}] \cdot [\text{DMn}] + k_{off(\text{CaDMn})} \cdot [\text{CaDMn}] \\ & - k_{on(\text{MgDMn})} \cdot [\text{Mg}^{2+}] \cdot [\text{DMn}] + k_{off(\text{MgDMn})} \cdot [\text{MgDMn}] - J_{fast(\text{DMn})} - J_{slow(\text{DMn})} \end{aligned} \quad (21)$$

and

$$J_{fast(\text{DMn})} = \frac{d[\text{DMn}]_f}{dt} = \frac{[\text{DMn}]_f}{\tau_{fast}} \cdot T_{flash} \quad (22)$$

$$J_{slow(\text{DMn})} = \frac{d[\text{DMn}]_s}{dt} = \frac{[\text{DMn}]_s}{\tau_{slow}} \cdot T_{flash} \quad (23)$$

$$J_{fast(\text{CaDMn})} = \frac{d[\text{CaDMn}]_f}{dt} = \frac{[\text{CaDMn}]_f}{\tau_{fast}} \cdot T_{flash} \quad (24)$$

$$J_{slow(\text{CaDMn})} = \frac{d[\text{CaDMn}]_s}{dt} = \frac{[\text{CaDMn}]_s}{\tau_{slow}} \cdot T_{flash} \quad (25)$$

$$J_{fast(\text{MgDMn})} = \frac{d[\text{MgDMn}]_f}{dt} = \frac{[\text{MgDMn}]_f}{\tau_{fast}} \cdot T_{flash} \quad (26)$$

$$J_{slow(\text{MgDMn})} = \frac{d[\text{MgDMn}]_s}{dt} = \frac{[\text{MgDMn}]_s}{\tau_{slow}} \cdot T_{flash} \quad (27)$$

$$\text{where } T_{flash} = 0 \text{ if } t < t_{flash} \quad (28)$$

$$\text{and } T_{flash} = 1 \text{ if } t \geq t_{flash} \quad (29)$$

For the photoproducts:

$$\frac{d[CaPP1]}{dt} = k_{on(CaDMn)} \cdot [Ca^{2+}] \cdot [PP1] - k_{off(CaPP1)} \cdot [CaPP1] + J_{fast(CaDMn)} + J_{slow(CaDMn)} \quad (30)$$

$$\frac{d[CaPP2]}{dt} = k_{on(CaDMn)} \cdot [Ca^{2+}] \cdot [PP2] - k_{off(CaPP2)} \cdot [CaPP2] \quad (31)$$

$$\frac{d[MgPP1]}{dt} = k_{on(MgDMn)} \cdot [Mg^{2+}] \cdot [PP1] - k_{off(MgPP)} \cdot [MgPP1] \quad (32)$$

$$\frac{d[MgPP2]}{dt} = k_{on(MgDMn)} \cdot [Mg^{2+}] \cdot [PP2] - k_{off(MgPP)} \cdot [CaPP2] + J_{fast(MgDMn)} + J_{slow(MgDMn)} \quad (33)$$

$$\begin{aligned} \frac{d[PP1]}{dt} = & -k_{on(CaDMn)} \cdot [Ca^{2+}] \cdot [PP1] + k_{off(CaPP1)} \cdot [CaPP1] \\ & - k_{on(MgDMn)} \cdot [Mg^{2+}] \cdot [PP1] + k_{off(MgPP)} \cdot [MgPP1] \\ & + J_{fast(CaDMn)} + J_{slow(CaDMn)} \end{aligned} \quad (34)$$

$$\begin{aligned} \frac{d[PP2]}{dt} = & -k_{on(CaDMn)} \cdot [Ca^{2+}] \cdot [PP2] + k_{off(CaPP2)} \cdot [CaPP2] \\ & - k_{on(MgDMn)} \cdot [Mg^{2+}] \cdot [PP2] + k_{off(MgPP)} \cdot [MgPP2] \\ & + 2 \cdot J_{fast(DMn)} + 2 \cdot J_{slow(DMn)} + J_{fast(MgDMn)} + J_{slow(MgDMn)} \end{aligned} \quad (35)$$

For the dye

$$\frac{d[CaD]}{dt} = k_{on(D)} \cdot [Ca^{2+}] \cdot [D] - k_{off(D)} \cdot [CaD] \quad (36)$$

$$[D] = [D]_T - [CaD] \quad (37)$$

For Ca^{2+} and Mg^{2+}

$$[Ca^{2+}] = [Ca^{2+}]_T - [CaDMn] - [CaPP1] - [CaPP2] \quad (38)$$

$$[Ca^{2+}]_T = [Ca^{2+}]_{t=0} + [CaDMn]_{t=0} \quad (39)$$

$$[Mg^{2+}] = [Mg^{2+}]_T - [MgDMn] - [MgPP1] - [MgPP2] \quad (40)$$

Initial conditions

$$[DMn]_{t=0} = [DMn]_T - [CaDMn]_{t=0} - [MgDMn]_{t=0} \quad (41)$$

$$[CaDMn]_{t=0} = \frac{K_{d(MgDMn)} \cdot [Ca^{2+}]_{t=0}}{K_{d(MgDMn)} \cdot [Ca^{2+}]_{t=0} + K_{d(CaDMn)} \cdot [Mg^{2+}]_{t=0} + K_{d(CaDMn)} \cdot K_{d(MgDMn)}} \cdot [DMn]_T \quad (42)$$

$$[MgDMn]_{t=0} = \frac{K_{d(CaDMn)} \cdot [Mg^{2+}]_{t=0}}{K_{d(CaDMn)} \cdot [Mg^{2+}]_{t=0} + K_{d(MgDMn)} \cdot [Ca^{2+}]_{t=0} + K_{d(CaDMn)} \cdot K_{d(MgDMn)}} \cdot [DMn]_T \quad (43)$$

$$[Mg^{2+}] = \frac{K_{d(CaDMn)} \cdot [Mg^{2+}]_T - K_{d(MgDMn)} \left(K_{d(CaDMn)} - [Ca^{2+}] - [DMn]_T \right) + X}{2K_{d(CaDMn)}} \quad (44a)$$

$$X = \sqrt{K_{d(CaDMn)}^2 \left([Mg^{2+}]_T \left([Mg^{2+}]_T + 2 \cdot K_{d(MgDMn)} - 2 \cdot [DMn]_T \right) + [DMn]_T \left([DMn]_T + 2 \cdot K_{d(MgDMn)} \right) + K_{d(MgDMn)}^2 \right) + K_{d(MgDMn)} \cdot [Ca^{2+}] \left(K_{d(MgDMn)} \cdot [Ca^{2+}] + 2 \cdot K_{d(CaDMn)} \left([Mg^{2+}]_T + K_{d(MgDMn)} [DMn]_T \right) \right)} \quad (44b)$$

(also see appendix A, see online journal)

Initially there are no photoproducts

$$[CaD]_{t=0} = \frac{[Ca^{2+}]_{t=0}}{K_{d(D)} + [Ca^{2+}]_{t=0}} \cdot [D]_T \quad (45)$$

Model output

$$\frac{\Delta F}{F} = \frac{\Delta F(t)}{F_{t=0}} = \frac{[CaD] \cdot (F_{ratio} - 1) + [D]_T}{[CaD]_{t=0} \cdot (F_{ratio} - 1) + [D]_T} \quad (46)$$